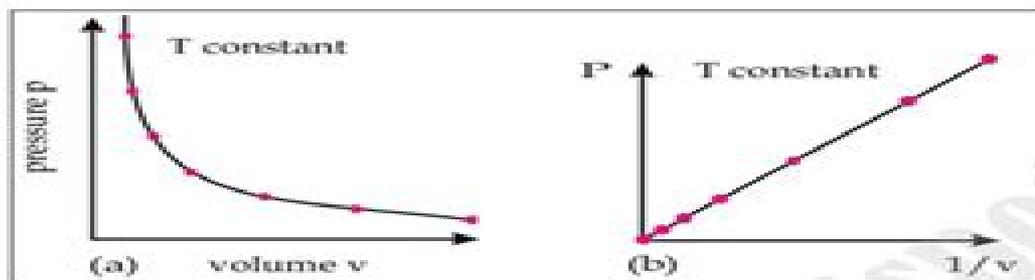


- (1) If the pressure of the gas is double volume becomes half. If pressure increases by three times then the volume becomes one-third and so on.
- (2) The graph between "p" and "V" between "p and $\frac{1}{V}$ " is shown below:



- (3) The graph between "p" and "V" shows that if pressure increases then volume decreases and vice-versa.
- (4) The graph between "p" and " $1/V$ " shows a straight line passing through the origin.
- (5) At constant temperature the product of pressure and volume is constant. i.e.
 $pV = \text{constant}$
- (6) Using the above result, at a constant temperature, we can write:
 initial pressure \times initial volume = final pressure \times final volume
 $p_1V_1 = p_2V_2$

Thus Robert Boyle concluded his Law known as "Boyle's Law" which states that:

Statement of Boyle's Law:

The volume of a fixed mass of a gas is inversely proportional to its pressure, provided its temperature remains constant.

Q.23: Write down the application of Boyle's Law.

Ans: Applications of (p-V) relationship of gas / Boyle's Law:

Some applications of pressure-volume (p-V) relationship of a gas i.e. Boyle's Law is given below:

Aerosols, such as spray paints, use Boyle's Law in their working mechanism.

When we breathe, our diaphragm moves downward, increasing the volume of the Lungs. This causes the pressure inside the Lungs to be less than the outside pressure so air rushes in.

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Table of Contents Physics Chapter Properties Of Matter Answers

1. Understanding the eBook Physics Chapter Properties Of Matter Answers
 - The Rise of Digital Reading Physics Chapter Properties Of Matter Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physics Chapter Properties Of Matter Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Chapter Properties Of Matter Answers
 - Personalized Recommendations
 - Physics Chapter Properties Of Matter Answers User Reviews and Ratings
 - Physics Chapter Properties Of Matter Answers and Bestseller Lists
5. Accessing Physics Chapter Properties Of Matter Answers Free and Paid eBooks
 - Physics Chapter Properties Of Matter Answers Public Domain eBooks
 - Physics Chapter Properties Of Matter Answers eBook Subscription Services
 - Physics Chapter Properties Of Matter Answers Budget-Friendly Options
6. Navigating Physics Chapter Properties Of Matter Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Physics Chapter Properties Of Matter Answers Compatibility with Devices
 - Physics Chapter Properties Of Matter Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physics Chapter Properties Of Matter Answers
 - Highlighting and Note-Taking Physics Chapter Properties Of Matter Answers
 - Interactive Elements Physics Chapter Properties Of Matter Answers

8. Staying Engaged with Physics Chapter Properties Of Matter Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physics Chapter Properties Of Matter Answers
9. Balancing eBooks and Physical Books Physics Chapter Properties Of Matter Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics Chapter Properties Of Matter Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physics Chapter Properties Of Matter Answers
 - Setting Reading Goals Physics Chapter Properties Of Matter Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physics Chapter Properties Of Matter Answers
 - Fact-Checking eBook Content of Physics Chapter Properties Of Matter Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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