

Section 1.1

Section 1.1

Section 1.1 covers the following topics in the areas provided:

1. The structure of a chemical compound depends on:
 - (a) the number of atoms of each element and protons
 - (b) the relative number of neutrons of isotopes and protons
 - (c) number of atoms of each element in each compound
 - (d) number of valence electrons, electronegativity, etc.
2. The structure of a chemical compound depends on:
 - (a) the number of atoms of each element and protons
 - (b) the number of valence electrons of each element and protons
 - (c) the number of neutrons of isotopes that combine with a certain quantity of protons
 - (d) the relative electronegativity of each element (e.g., C, O, N, H, etc.)
3. The structure of a chemical compound depends on:
 - (a) the number of atoms of each element and protons
 - (b) the energy released in the reaction
 - (c) relative electronegativity of all elements in the reaction
 - (d) number of valence electrons of each element
4. The structure of a chemical compound depends on:
 - (a) the number of atoms of each element and protons
 - (b) the number of neutrons
 - (c) the number of protons
 - (d) relative electronegativity

Section 1.1 covers the following topics in the areas provided:

$$2H_2 + O_2 \rightarrow 2H_2O$$

a. Balance the equation:



Given 100 g of H_2 and 100 g of O_2 , what mass of H_2O is formed?

1. The mass of H_2O formed is 100 g. (a) 100 g (b) 200 g (c) 300 g (d) 400 g

2. The mass of H_2O formed is 100 g. (a) 100 g (b) 200 g (c) 300 g (d) 400 g

3. The mass of H_2O formed is 100 g. (a) 100 g (b) 200 g (c) 300 g (d) 400 g

Chapter 9 Review Stoichiometry Section 3

Siddappa N.Byrareddy



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engineering *Foundations of College Chemistry* Morris Hein, Susan Arena, 1993 This is the Student Study Guide to accompany Hein's *Foundations of College Chemistry Alternate Edition 12th Edition* Prentice Hall Chemistry ,2000 5

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