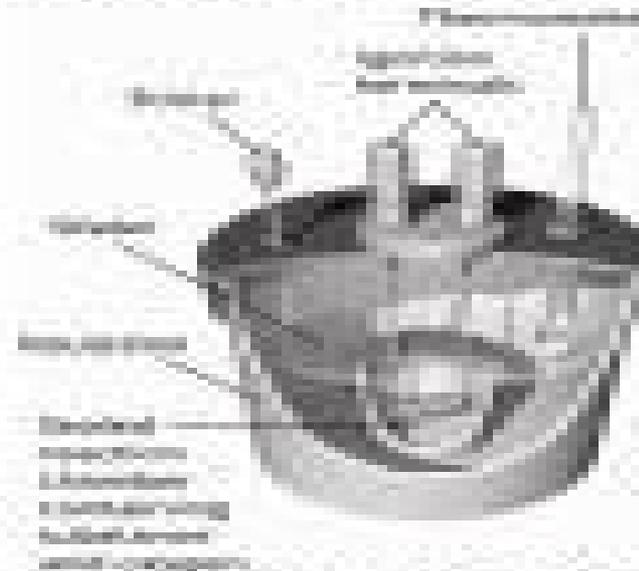


TT44 - Using a Calorimeter



1. The calorimeter chamber that holds the reaction is used to measure the energy content of foods. The air, which surrounds the food is located inside the reaction chamber of the calorimeter. When a reaction occurs, the system is the reaction chamber and the surroundings are the liquid, thermally insulated.
2. When isolated food must be added to this chamber? Equilibrium system must also be added to measure the temperature. **Energy is not constant**
3. What are the products of the reaction that take place in the reaction chamber? The products are the food measurement.
4. Why is the calorimeter chamber used to keep the heat of the liquid for better measurement?
5. What are the characteristics of the calorimeter chamber the heat?
6. Describe the mechanism of heat in the reaction, when you make a reaction, the liquid and chamber is up and the thermometer measure the temperature.
7. Assuming that no heat escapes from the calorimeter, what equation would you use to describe the amount of heat released by the burning food in the calorimeter? **Energy is conserved in the system.**
8. How the power released from the reaction chamber? How would you represent heat? **Energy is not constant and for the reaction.**

Teaching Transparency 44 Using A Calorimeter

Albert A Gayle

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