

# AP Chemistry - Chemical Kinetics

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

- After five half-life periods for a first order reaction, what fraction of reactant remains?
- If you plot  $1/[\text{reactant}]$  versus time and observe a straight line, what is the order of the reaction? If  $\ln[\text{reactant}]$  is plotted versus time, and a straight line of negative slope is observed, what is the order of the reaction?
- Indicate whether each of the following statements is true or false. Change the wording of each false statement to make it true.
  - In stepwise reactions the rate-determining step is the slow one.
  - It is possible to change the rate constant for a reaction by changing the temperature.
  - As a first order reaction proceeds at a constant temperature, the rate remains constant.
  - The rate constant for a reaction is independent of reactant concentrations.
- For each reaction below give the relative rates of formation of products and disappearance of reactants using coefficients and rate change symbols.
  - $2\text{O}_2(\text{g}) \rightarrow 3\text{O}_2(\text{g})$
  - $2\text{HOF}(\text{g}) \rightarrow 2\text{HF}(\text{g}) + \text{O}_2(\text{g})$
  - $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$
- The reaction  $\text{CO}(\text{g}) + \text{NO}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{NO}(\text{g})$  is second order in  $\text{NO}_2$  and zero order in  $\text{CO}$  at temperatures less than 500K.
  - Write the rate law for the reaction.
  - How will the reaction rate change if the  $\text{NO}_2$  concentration is halved?
- For a reaction of the type  $\text{A} + \text{B} \rightarrow \text{C}$ , it is experimentally observed that doubling the concentration of B causes the reaction rate to be increased fourfold, but doubling the concentration of A has no effect on the rate. Determine the initial rate equation.
- A study of the reaction  $2\text{A} + \text{B} \rightarrow \text{C} + \text{D}$  gave the following results:

Exprmnt #	Init Con (M)		Init Rate (M/s)
	[A] <sub>0</sub>	[B] <sub>0</sub>	
1.	0.10	0.05	$6.0 \times 10^{-3}$
2.	0.20	0.05	$1.2 \times 10^{-2}$
3.	0.30	0.05	$1.8 \times 10^{-2}$
4.	0.20	0.15	$1.1 \times 10^{-1}$

- What is the rate law for this reaction?
  - What is the overall order of the reaction?
  - Could this reaction occur in a single step?
- The decomposition of  $\text{N}_2\text{O}_5$  (to give  $\text{NO}_2$  and  $\text{O}_2$ ) follows the rate law "rate =  $k[\text{N}_2\text{O}_5]^n$ " where  $k$  is  $5.0 \times 10^{-4}$  per second at a particular temperature.
    - What is the half-life of this reaction?
    - How long does it take for the  $\text{N}_2\text{O}_5$  concentration to drop to one tenth of its original value?

# Ap Chemistry Chemical Kinetics Worksheet Answers

**M Planty**



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This book delves into Ap Chemistry Chemical Kinetics Worksheet Answers. Ap Chemistry Chemical Kinetics Worksheet Answers is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Ap Chemistry Chemical Kinetics Worksheet Answers, encompassing both the fundamentals and more intricate discussions.

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