

Chapter 8 Chemical Reactions Chapter Review Answers

1. The evolution of energy, the production of a gas, the formation of a precipitate, and a color change.
2. A correctly written chemical equation must represent the known facts, contain the correct formulas for the reactants and products, and satisfy the law of conservation of mass.
3.
 - a. It is a number that appears in front of a formula in a chemical equation.
 - b. The number of atoms of each type in the formula is multiplied by the coefficient.
4. methane + oxygen \rightarrow carbon dioxide + water;
 $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$;
 $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$
5. A chemical equation provides the number of molecules, moles, and atoms of each element or compound composing the reactants and products as well as the mole ratios of the reactants to each other and to the products.
6. Word and formula equations have only qualitative meaning; they do not express the relative quantities of a reaction's reactants or products.
7.
 - a. a solution in water
 - b. a substance that accelerates a chemical reaction but can be recovered unchanged
 - c. a chemical reaction in which the products re-form the original reactants
8.
 - a. KOH
 - b. $\text{Ca}(\text{NO}_3)_2$
 - c. Na_2CO_3
 - d. CCl_4
 - e. MgBr_2
9. (1) Balance different types of atoms one at a time; (2) balance types of atoms that appear only once on each side of the equation first; (3) balance as single units any polyatomic ions that appear on both sides of the equation; and (4) balance H atoms and O atoms last.
10.
 - a. 6N
 - b. 4H, 2O
 - c. 4H, 4N, 12O
 - d. 2Ca, 4O, 4H
 - e. 3Ba, 6Cl, 18O
 - f. 5Fe, 10N, 30O
 - g. 12Mg, 8P, 32O
 - h. 4N, 16H, 2S, 8O
 - i. 12Al, 18Se, 72O
 - j. 12C, 32H
11.
 - a. $2\text{ZnS}(\text{s}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{ZnO}(\text{s}) + 2\text{SO}_2(\text{g})$
 - b. $2\text{HCl}(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow \text{BaCl}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
 - c. $2\text{HNO}_3(\text{aq}) + \text{Ca}(\text{OH})_2(\text{aq}) \rightarrow \text{Ca}(\text{NO}_3)_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
12.
 - a. Solid zinc sulfide reacts with oxygen gas to form solid zinc oxide and sulfur dioxide gas.
 - b. When solid calcium hydride is added to water, aqueous calcium hydroxide and hydrogen gas are formed.
 - c. Aqueous silver nitrate mixed with aqueous potassium iodide produces a precipitate of silver iodide and aqueous potassium nitrate.
13.
 - a. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - b. $2\text{Al} + \text{Fe}_2\text{O}_3 \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$
 - c. $\text{Pb}(\text{CH}_3\text{COO})_2 + \text{H}_2\text{S} \rightarrow \text{PbS} + 2\text{CH}_3\text{COOH}$
14.
 - a. LiO_2 is an incorrect formula; $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$
 - b. H_2Cl_2 is an incorrect formula; $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - c. MgO_2 is an incorrect formula, and the equation as written is not balanced; $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
 - d. I is an incorrect formula for iodine, and the equation is not balanced; $2\text{NaI} + \text{Cl}_2 \rightarrow 2\text{NaCl} + \text{I}_2$
15.
 - a. $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 - b. $\text{P}_2\text{O}_5 + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$
 - c. $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
16. $\text{CH}_4(\text{g}) + 4\text{Cl}_2(\text{g}) \rightarrow \text{CCl}_4(\text{l}) + 4\text{HCl}(\text{g})$
17.
 - a. $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
 - b. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 - c. $2\text{Li} + \text{Cl}_2 \rightarrow 2\text{LiCl}$
 - d. $\text{Ca} + \text{I}_2 \rightarrow \text{CaI}_2$
18.
 - (1) In a synthesis reaction, two or more substances combine to form a new substance: $\text{A} + \text{X} \rightarrow \text{AX}$.
 - (2) In a decomposition reaction, a single compound undergoes a reaction that produces two or more simpler substances: $\text{AX} \rightarrow \text{A} + \text{X}$.
 - (3) In a single-displacement reaction, one element replaces another element in a compound:
 $\text{A} + \text{BX} \rightarrow \text{AX} + \text{B}$ or $\text{Y} + \text{BX} \rightarrow \text{BY} + \text{X}$

Chapter 8 Chemical Reactions Answers

Wenbin Ji



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Chemistry (Teacher Guide) Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books. The teacher is one who guides students through the subject matter, helps each student stay on schedule and be organized, and is their source of accountability along the way. With that in mind, this guide provides additional help through the laboratory exercises as well as lessons, quizzes, and examinations that are provided along with the answers. The lessons in this study emphasize working through procedures and problem solving by learning patterns. The vocabulary is kept at the essential level. Practice exercises are given with their answers so that the patterns can be used in problem solving. These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college. Guided labs are provided to enhance instruction of weekly lessons. There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions. It is important to see the hand of God and His principles and wisdom as it plays out in chemistry. This course integrates what God has told us in the context of this study. Features: Each suggested weekly schedule has five easy-to-manage lessons that combine reading and worksheets. Worksheets, quizzes, and tests are perforated and three-hole punched; materials are easy to tear out, hand out, grade, and store. Adjust the schedule and materials needed to best work within your educational program. Space is given for assignments, dates, and notes. There is flexibility in scheduling. Adapt the days to your school schedule. Workflow: Students will read the pages in their book and then complete each section of the teacher guide. They should be encouraged to complete as many of the activities and projects as possible as well. Tests are given at regular intervals with space to record each grade. About the Author: DR DENNIS ENGLIN earned his bachelor's from Westmont College, his master of science from California State University, and his EdD from the University of Southern California. He enjoys teaching animal biology, vertebrate biology, wildlife biology, organismic biology, and astronomy at The Master's University. His professional memberships include the Creation Research Society, the American Fisheries Association, Southern California Academy of Sciences, Yellowstone Association, and Au Sable Institute of Environmental Studies.

Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts. After a brief review of each section's critical ideas, students are taken through stepped-out worked examples, try-it-yourself examples, and chapter quizzes, all structured to reinforce chapter objectives and build problem-solving techniques. The solutions manual includes detailed solutions to all odd-numbered exercises in the text.

Essentials of Chemical Reaction Engineering H. Scott Fogler, 2011 Accompanying DVD-ROM contains many realistic interactive simulations.

Study Guide with Answers to Selected Problems H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview, activities, and practice

problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

Stochastic and Dynamic Views of Chemical Reaction Kinetics in Solutions A. M. Kuznetsov (dokt. fiz.-mat. nauk.), Alexander M. Kuznetsov, 1999-01-01 *Organic Chemistry, 5e Student Study Guide and Solutions Manual*

David R. Klein, Laurie S. Starkey, 2025-03-04 Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems With *Organic Chemistry Student Study Guide and Solutions Manual 5th Edition* students can learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry

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General Organic and Biological Chemistry Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features

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