

Key.

- b) 6.45 moles Cr_2O_3 $\text{FW} = 152.0 \text{ g/mol}$ 980 g
 $6.45 \text{ mol} \times \left(\frac{152.0 \text{ g}}{1 \text{ mol}} \right) = 980 \text{ g}$
- c) 0.144 mole K_2S $\text{FW} = 110.3 \text{ g/mol}$ 15.9 g
 $0.144 \text{ mol} \times \left(\frac{110.3 \text{ g}}{1 \text{ mol}} \right) = 15.9 \text{ g}$
- d) 5.0 moles $\text{C}_2\text{H}_2\text{O}_2$ 180.0 g/mol 900 g = 9.0 \times 10^2 g
 $5.0 \text{ mol} \times \left(\frac{180.0 \text{ g}}{1 \text{ mol}} \right) = 900 \text{ g}$
- e) 5.10×10^{-4} mole NaCl $\text{FW} = 58.5 \text{ g/mol}$ 0.0298 g = 2.98 \times 10^{-2} g
 $5.10 \times 10^{-4} \text{ mol} \times \left(\frac{58.5 \text{ g}}{1 \text{ mol}} \right) = 0.0298 \text{ g}$

Solve the following stoichiometry problems. Write a balanced equation for each.

- a) How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas, to produce hydrogen chloride gas? Equation: $\text{H}_2 + \text{Cl}_2 \rightarrow 2 \text{HCl}$
 $15.1 \text{ g} \left(\frac{1 \text{ mol}}{71.0 \text{ g}} \right) = 0.213 \text{ mol Cl}_2$
 $0.213 \text{ mol Cl}_2 \left(\frac{1 \text{ mol H}_2}{1 \text{ mol Cl}_2} \right) = 0.213 \text{ mol H}_2$ 0.213 mol H₂
- b) How many grams of oxygen gas are needed to react completely with 16.2 g of hydrogen gas, to produce water? Equation: $2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$
 $16.2 \text{ g} \left(\frac{1 \text{ mol}}{2.0 \text{ g}} \right) = 8.10 \text{ mol H}_2$
 $8.10 \text{ mol H}_2 \left(\frac{1 \text{ mol O}_2}{2 \text{ mol H}_2} \right) = 4.05 \text{ mol O}_2$ 130 g O₂
 $4.05 \text{ mol O}_2 \left(\frac{32.0 \text{ g}}{1 \text{ mol}} \right) = 130 \text{ g O}_2$
- c) How many moles of calcium oxide are produced when 36.5 g of calcium react completely with oxygen gas? Equation: $2 \text{Ca} + \text{O}_2 \rightarrow 2 \text{CaO}$
 $36.5 \text{ g} \left(\frac{1 \text{ mol}}{40.1 \text{ g}} \right) = 0.910 \text{ mol Ca}$
 $0.910 \text{ mol Ca} \left(\frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} \right) = 0.910 \text{ mol CaO}$ 0.910 mol CaO
- d) How many grams of aluminum are required to produce 410g of aluminum oxide, through reaction with oxygen gas? Equation: $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
 $410 \text{ g} \left(\frac{1 \text{ mol}}{102.0 \text{ g}} \right) = 4.02 \text{ mol Al}_2\text{O}_3$
 $4.02 \text{ mol Al}_2\text{O}_3 \left(\frac{4 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \right) = 8.04 \text{ mol Al}$ 217 g = 220 g (2 sf)
 $8.04 \text{ mol Al} \left(\frac{27 \text{ g}}{1 \text{ mol}} \right) = 217 \text{ g}$
- e) How many grams of solid iodine are produced by the decomposition of 25.4 grams of hydrogen iodide gas? Equation: $2 \text{HI} \rightarrow \text{H}_2 + \text{I}_2$
 $25.4 \text{ g} \left(\frac{1 \text{ mol}}{127.9 \text{ g}} \right) = 0.199 \text{ mol HI}$
 $0.199 \text{ mol HI} \left(\frac{1 \text{ mol I}_2}{2 \text{ mol HI}} \right) = 0.0993 \text{ mol I}_2$ 25.2 g I₂
 $0.0993 \text{ mol I}_2 \left(\frac{253.8 \text{ g}}{1 \text{ mol}} \right) = 25.2 \text{ g I}_2$

Chapter 9 Review Stoichiometry Section 1 Answer Key

Minjie Lin



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Table of Contents Chapter 9 Review Stoichiometry Section 1 Answer Key

1. Understanding the eBook Chapter 9 Review Stoichiometry Section 1 Answer Key
 - The Rise of Digital Reading Chapter 9 Review Stoichiometry Section 1 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Review Stoichiometry Section 1 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 9 Review Stoichiometry Section 1 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Review Stoichiometry Section 1 Answer Key
 - Personalized Recommendations

- Chapter 9 Review Stoichiometry Section 1 Answer Key User Reviews and Ratings
- Chapter 9 Review Stoichiometry Section 1 Answer Key and Bestseller Lists
- 5. Accessing Chapter 9 Review Stoichiometry Section 1 Answer Key Free and Paid eBooks
 - Chapter 9 Review Stoichiometry Section 1 Answer Key Public Domain eBooks
 - Chapter 9 Review Stoichiometry Section 1 Answer Key eBook Subscription Services
 - Chapter 9 Review Stoichiometry Section 1 Answer Key Budget-Friendly Options
- 6. Navigating Chapter 9 Review Stoichiometry Section 1 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 9 Review Stoichiometry Section 1 Answer Key Compatibility with Devices
 - Chapter 9 Review Stoichiometry Section 1 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 9 Review Stoichiometry Section 1 Answer Key
 - Highlighting and Note-Taking Chapter 9 Review Stoichiometry Section 1 Answer Key
 - Interactive Elements Chapter 9 Review Stoichiometry Section 1 Answer Key
- 8. Staying Engaged with Chapter 9 Review Stoichiometry Section 1 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 9 Review Stoichiometry Section 1 Answer Key
- 9. Balancing eBooks and Physical Books Chapter 9 Review Stoichiometry Section 1 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 9 Review Stoichiometry Section 1 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 9 Review Stoichiometry Section 1 Answer Key
 - Setting Reading Goals Chapter 9 Review Stoichiometry Section 1 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 9 Review Stoichiometry Section 1 Answer Key
 - Fact-Checking eBook Content of Chapter 9 Review Stoichiometry Section 1 Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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