

Section 10.1

Section 10.1

Answer complete answer multiple-choice questions that are provided.

1. _____
 - a. _____
 - (a) amount, in grams, of all the reactants and products.
 - (b) volume of molecular oxygen of reactants and products.
 - (c) number of moles of work done by work done by reactants.
 - (d) number of atoms of each element in the reactants.
 - b. _____
 - (a) the mass number of all the reactants and products.
 - (b) the mass of carbon produced when a certain mass of oxygen is consumed.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the amount of energy represented by the heat of combustion.
 - c. _____
 - (a) number of molecules of each element in the reactants and products.
 - (b) number of moles of each element in the reactants.
 - (c) volume of molecular oxygen of reactants and products.
 - (d) number of molecules of each element in the reactants.
 - d. _____
 - (a) number of molecules of each element in the reactants and products.
 - (b) number of moles of each element in the reactants.
 - (c) number of molecules of each element in the reactants.
 - (d) number of molecules of each element in the reactants.
2. _____
 - a. _____
 - (a) number of molecules of each element in the reactants and products.
 - (b) number of moles of each element in the reactants.
 - (c) volume of molecular oxygen of reactants and products.
 - (d) number of molecules of each element in the reactants.
 - b. _____
 - (a) number of molecules of each element in the reactants and products.
 - (b) number of moles of each element in the reactants.
 - (c) volume of molecular oxygen of reactants and products.
 - (d) number of molecules of each element in the reactants.
 - c. _____
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 - (b) number of moles of each element in the reactants.
 - (c) volume of molecular oxygen of reactants and products.
 - (d) number of molecules of each element in the reactants.
 - d. _____
 - (a) number of molecules of each element in the reactants and products.
 - (b) number of moles of each element in the reactants.
 - (c) volume of molecular oxygen of reactants and products.
 - (d) number of molecules of each element in the reactants.

3. Answer the two questions based on the following balanced equation:



a. Balance the equation:



b. What is the coefficient of H_2 in H_2O ?

c. How many moles of H_2 are required to produce 1.0 mole of H_2O ?

d. How many moles of O_2 are required to produce 1.0 mole of H_2O ?

e. How many moles of H_2 are required to produce 1.0 mole of H_2O ?

Chapter 9 Review Stoichiometry Section 1 Answers

Morris Hein



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