

2018

Name: Phyllis, J. Smith Lab Section: 201

Suppose that you had the correct apparatus for the following experiment. In what state would the liquid be at room temperature?

Compound	mp, °C	bp, °C
<u>1,2-dichloroethane</u>	<u>-100</u>	<u>84</u>
<u>1,1-dichloroethane</u>	<u>-100</u>	<u>59</u>
<u>1,1,1-trichloroethane</u>	<u>-110</u>	<u>74</u>

1. If you had three moles of each of the 1,1-dichloroethane, 1,1,1-trichloroethane, and 1,2-dichloroethane, how many moles of water would be produced by combustion of each to yield  $\text{CO}_2$  and  $\text{H}_2\text{O}$ ?

$$\begin{array}{l}
 1,2\text{-dichloroethane} \quad 3 \text{ moles} \quad \left| \quad \begin{array}{l} \text{C}_2\text{H}_4\text{Cl}_2 \\ 98.96 \text{ g/mol} \end{array} \right| \quad \begin{array}{l} \text{C}_2\text{H}_6\text{Cl}_2 \\ 147.03 \text{ g/mol} \end{array} \quad \left| \quad \begin{array}{l} \text{C}_2\text{H}_5\text{Cl}_3 \\ 203.03 \text{ g/mol} \end{array} \right| \quad \begin{array}{l} \text{H}_2\text{O} \\ 18.015 \text{ g/mol} \end{array} \\
 \hline
 \text{moles of H}_2\text{O} = 6 \text{ moles}
 \end{array}$$

2. If you had 100 g of the mixture and density of the mixture is 1.20 g/cm<sup>3</sup>, how many moles of water would be produced by combustion of the mixture?

$$\begin{array}{l}
 100 \text{ g} \quad \left| \quad \begin{array}{l} 100 \text{ g} \\ 1.20 \text{ g/cm}^3 \end{array} \right| \quad \begin{array}{l} 83.3 \text{ cm}^3 \\ 244.2 \text{ g} \end{array} \quad \left| \quad \begin{array}{l} 1.20 \text{ g/cm}^3 \\ 1.20 \text{ g/cm}^3 \end{array} \right| \quad \begin{array}{l} 100 \text{ g} \\ 100 \text{ g} \end{array} \\
 \hline
 \text{moles of H}_2\text{O} = 27.17 \text{ moles}
 \end{array}$$

$$\begin{array}{l}
 100 \text{ g} \quad \left| \quad \begin{array}{l} 100 \text{ g} \\ 1.20 \text{ g/cm}^3 \end{array} \right| \quad \begin{array}{l} 83.3 \text{ cm}^3 \\ 244.2 \text{ g} \end{array} \quad \left| \quad \begin{array}{l} 1.20 \text{ g/cm}^3 \\ 1.20 \text{ g/cm}^3 \end{array} \right| \quad \begin{array}{l} 100 \text{ g} \\ 100 \text{ g} \end{array} \\
 \hline
 \text{moles of H}_2\text{O} = 27.17 \text{ moles}
 \end{array}$$

27.17 moles  
27.17 moles  
27.17 moles

# Chem 110 Lab Questions And Answers

**Ying-Ying Zheng**



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