

# Experiment 18-Chemical Reactions Lab Report

Name: \_\_\_\_\_

Section: \_\_\_\_\_

For this lab, you need to balance each equation, and classify the reaction according to the types of reactions listed in the lab handout using C (combination), D (decomposition), Cu (combustion), S (single replacement) and DR (double replacement). Include what you observed below each reaction.

- |  | <u>Type of Reaction</u> |
|--|-------------------------|
| 1. $\underline{\hspace{1cm}} \text{CaCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{Na}_3\text{PO}_4 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{Ca}_3(\text{PO}_4)_2 (\text{s}) + \underline{\hspace{1cm}} \text{NaCl} (\text{aq})$  | _____                   |
| Observations:  |                         |
| 2. $\underline{\hspace{1cm}} \text{CuSO}_4 (\text{H}_2\text{O})_5 + \text{heat} \rightarrow \underline{\hspace{1cm}} \text{CuSO}_4 (\text{s}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{g})$  | _____                   |
| Observations:  |                         |
| $\underline{\hspace{1cm}} \text{CuSO}_4 (\text{s}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l}) \rightarrow \underline{\hspace{1cm}} \text{CuSO}_4 (\text{H}_2\text{O})_5 (\text{s})$  | _____                   |
| Observations:  |                         |
| 3. $\underline{\hspace{1cm}} \text{Zn} (\text{s}) + \underline{\hspace{1cm}} \text{HCl} (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{ZnCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{H}_2 (\text{g})$   | _____                   |
| Observations:  |                         |
| $\underline{\hspace{1cm}} \text{H}_2 (\text{g}) + \underline{\hspace{1cm}} \text{O}_2 (\text{g}) \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{g})$ The "popping" reaction induced with the match.  | _____                   |
| Observations:  |                         |
| 4. $\underline{\hspace{1cm}} \text{HCl} (\text{aq}) + \underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{CO}_2 (\text{g}) + \underline{\hspace{1cm}} \text{NaCl} (\text{aq}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l})$ | _____                   |
| Observations:  |                         |
| 5. $\underline{\hspace{1cm}} \text{H}_2\text{O}_2 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{O}_2 (\text{g}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l})$ (the $\text{MnO}_2$ is a catalyst, it does not react)   | _____                   |
| Observations:  |                         |
| 6. $\underline{\hspace{1cm}} \text{HCl} (\text{aq}) + \underline{\hspace{1cm}} \text{NaOH} (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{NaCl} (\text{aq}) + \underline{\hspace{1cm}} \text{H}_2\text{O} (\text{l})$  | _____                   |
| Observations:  |                         |
| 7. $\underline{\hspace{1cm}} \text{CaCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{NaCl} (\text{aq}) + \underline{\hspace{1cm}} \text{CaCO}_3 (\text{s})$   | _____                   |
| Observations:  |                         |
| 8. $\underline{\hspace{1cm}} \text{Zn} (\text{s}) + \underline{\hspace{1cm}} \text{CuSO}_4 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{ZnSO}_4 (\text{aq}) + \underline{\hspace{1cm}} \text{Cu} (\text{s})$   | _____                   |
| Observations:  |                         |
| 9. $\underline{\hspace{1cm}} \text{Cu} (\text{s}) + \underline{\hspace{1cm}} \text{ZnSO}_4 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{CuSO}_4 (\text{aq}) + \underline{\hspace{1cm}} \text{Zn} (\text{s})$   | _____                   |
| Observations:  |                         |
| 10. $\underline{\hspace{1cm}} \text{CaCl}_2 (\text{aq}) + \underline{\hspace{1cm}} \text{NaNO}_3 (\text{aq}) \rightarrow \underline{\hspace{1cm}} \text{Ca}(\text{NO}_3)_2 (\text{aq}) + \underline{\hspace{1cm}} \text{NaCl} (\text{aq})$   | _____                   |
| Observations:  |                         |

# Chemical Reactions Lab Answers

**S Baum**



## **Chemical Reactions Lab Answers:**

## **Chemical Reactions Lab Answers** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Chemical Reactions Lab Answers**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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