

Chem 1201 Final Exam

$$N_A = 6.022 \times 10^{23} \text{ things/mole}$$

$$1.6606 \times 10^{-24} \text{ g} = 1 \text{ amu}$$

- 1 Calculate the volume in liters of 82.0 m³ of natural gas. 1 cm³ = 1 mL.
- (A) 8.2 L
(B) 8.2×10^7 L
(C) 8.2×10^4 L
(D) 0.082 L
(E) 8.2×10^{-8} L
- 2 The molecule BrF₃ has a electron domain geometry and a molecular geometry.
- (A) tetrahedral, tetrahedral
(B) trigonal bipyramidal, seesaw
(C) tetrahedral, trigonal pyramidal
(D) trigonal bipyramidal, T-shaped
(E) trigonal planar, trigonal planar
- 3 What volume (mL) of 0.135 M sodium hydroxide is required to neutralize 440.0 mL of 0.584 M hydrochloric acid?
- (A) 5.25×10^{-3} mL
(B) 34.7 mL
(C) 1.9×10^3 mL
(D) 1.9 mL
(E) 6.47×10^3 mL
- 4 Which of these statements best describes the Heisenberg Uncertainty principle?
- (A) For each electron in an atom the magnetic quantum number m_l is uncertain.
(B) The probability of finding an electron in an atom is determined by ψ^2
(C) In the atomic world all measurable quantities are completely uncertain.

Chemistry 1 Final Exam Answers

Mark Steyvers



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