

## Chapter 13

# Hydrocarbons

### Question 13.1:

How do you account for the formation of ethane during chlorination of methane?

### Answer 13.1:

Chlorination of methane proceeds via a free radical chain mechanism. The whole reaction takes place in the given three steps.

#### Step 1: Initiation:

The reaction begins with the homolytic cleavage of Cl – Cl bond as:



#### Step 2: Propagation:

In the second step, chlorine free radicals attack methane molecules and break down the C–H bond to generate methyl radicals as:



These methyl radicals react with other chlorine free radicals to form methyl chloride along with the liberation of a chlorine free radical.

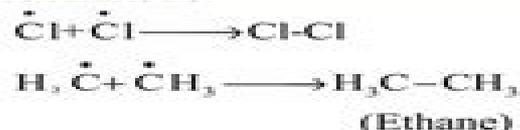


Hence, methyl free radicals and chlorine free radicals set up a chain reaction. While HCl and CH<sub>3</sub>Cl are the major products formed, other higher halogenated compounds are also formed as:



#### Step 3: Termination:

Formation of ethane is a result of the termination of chain reactions taking place as a result of the consumption of reactants as:



Hence, by this process, ethane is obtained as a by-product of chlorination of methane.

# Chemistry Chapter 13 Solutions Manual

**Tobias Bleicker**



**Chemistry Chapter 13 Solutions Manual:**

The book delves into Chemistry Chapter 13 Solutions Manual. Chemistry Chapter 13 Solutions Manual is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chemistry Chapter 13 Solutions Manual, encompassing both the fundamentals and more intricate discussions.

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- Chapter 2: Essential Elements of Chemistry Chapter 13 Solutions Manual
- Chapter 3: Chemistry Chapter 13 Solutions Manual in Everyday Life
- Chapter 4: Chemistry Chapter 13 Solutions Manual in Specific Contexts
- Chapter 5: Conclusion

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