

QUALITATIVE ANALYSIS

Experiment	Observation	Inference
<u>Preliminary Test</u>		
1. Noted the colour of the given salt	1. Blue or Bluish Green 2. Greenish 3. Light Green 4. Dark Brown 5. Light Pink 6. Colourless	Cu^{2+} or Ni^{2+} Ni^{2+} Fe^{2+} Fe^{3+} Mn^{2+} Absence of Cu^{+2} , Ni^{2+} , Fe^{2+} , Fe^{3+} , Mn^{2+}
2. Noted the solubility	1. Soluble in water 2. Soluble in dilute HCl	NH_4^+ may be present Absence of group I cations
3. Dry Heating Test: Heated a pinch of the salt in a dry Test tube	1. Colour, odourless gas 2. Colourless gas with rotten egg smell 3. Colourless gas with vinegar smell 4. Colourless gas with ammonia smell 5. Reddish brown gas 6. Greenish yellow gas 7. Violet gas	CO_3^{2-} may be present S^{2-} may be present CH_3COO^- may be present NH_4^+ may be present NO_3^- or Br^- may be present Cl^- may be present I^- may be present
<u>Identification of acid radical (Anion)</u>		
1. Dilute Acid Test: Treated a pinch of the substance With dilute H_2SO_4	1. Colourless, odourless gas with brisk effervescence which turned lime water milky. 2. Vinegar Smell 3. No characteristic Reaction	CO_3^{2-} present $\text{CO}_3^{2-} + 2\text{H}^+ \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ $\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$ CH_3COO^- Present $\text{CH}_3\text{COO}^- + \text{H}^+ \rightarrow \text{CH}_3\text{COOH}$ Absence of CO_3^{2-} , CH_3COO^-
2. Concentrated Acid Test: Treated a pinch of the substance With concentrated H_2SO_4	1. Colourless gas with pungent Smell which gave dense white fumes when a glass rod dipped in NH_4OH was brought near the mouth of the test tube 2. Reddish brown vapours with Pungent smell. Solution turns red 3. Deep violet vapours with pungent smell. Sides of test tube	Cl^- present $\text{Cl}^- + \text{H}^+ \rightarrow \text{HCl}$ $\text{HCl} + \text{NH}_4\text{OH} \rightarrow \text{NH}_4\text{Cl} + \text{H}_2\text{O}$ Br^- present $\text{Br}^- + \text{H}^+ \rightarrow \text{HBr}$ $2\text{HBr} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Br}_2 + \text{SO}_2$ I^- present $\text{I}^- + \text{H}^+ \rightarrow \text{HI}$ $2\text{HI} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{I}_2 + \text{SO}_2$

Chemistry Practical Qualitative Analysis Sheet

Frederick Hutton Getman



Chemistry Practical Qualitative Analysis Sheet:

Practical Chemistry. The Principles of Qualitative Analysis Sir William Augustus Tilden, 1880 A Laboratory Manual of Organic Chemistry for Beginners Arnold Frederick Holleman, 1913 **Practical Chemistry; the Principles of Qualitative Analysis** Sir William Augustus Tilden, 2013-09 This historic book may have numerous typos and missing text Purchasers can usually download a free scanned copy of the original book without typos from the publisher Not indexed Not illustrated 1886 edition Excerpt PRACTICAL CHEMISTRY PART I PRELIMINARY To make a Washing Bottle Take a clean widemouth bottle holding about a pint and fit a good sound cork tightly into the neck Then take a piece of stout glass tubing about inch external diameter and with the aid of a cork borer bore two holes through the cork parallel to each other and of such a size that the tube will pass through with some friction If you have not a cork borer of the exact size required take one a little smaller and then enlarge the hole by means of a round file Cut off a piece of the tube of about twice the height of the bottle Draw off one end to a point in the blowpipe flame Cut it off with a file and make the orifice smooth by holding it in the flame for a minute or two taking care not to close it completely Then in a common fish tail gas flame bend the tube so as to B form an angle of about 60 at such a distance from the wide end that when the tube is passed through the cork it will reach well to the bottom of the bottle Into the second hole in the cork another piece of the same tubing is fitted but this is cut off just below the cork and is bent to form an angle of about 120 The open end projecting upwards should be made smooth at the edge by holding it for a few minutes in quantities of water into test tubes and for general purposes as a store of water It is used by blowing gently into the upturned end of the mouthpiece and so producing a pressure upon the surface of the water in the bottle To Fix a Platinum Wire into a Glass Holder Take six inches of narrow glass tubing and by holding it in the Bunsen flame soften it so that it may be drawn out in the middle When cold scratch it at the middle with a file and break it The Annual Catalogue of Purdue University, Lafayette, Indiana ... with Announcements for ... Purdue University, 1920 **The Educational Times, and Journal of the College of Preceptors**, 1898 Elementary Linear Perspective of Form and Shadow Samuel Edward Warren, 1904 **The Medical Times and Gazette**, 1873 Annual Announcement of the School of Pharmacy Purdue University. School of Pharmacy, 1926 **Pharmaceutical Journal**, 1900 **Engineering**, 1906 **The pharmaceutical journal and transactions**, 1872 *Elementary instruction in chemical analysis [tr. from Anleitung zur qualitativen chemischen Analyse] ed. by J.L. Bullock* Carl Remigius Fresenius, 1872 *Dental Brief* Thomas Bromwell Welch, Wilbur F. Litch, Alfred P. Lee, 1907 **A systematic handbook of volumetric analysis or, The quantitative estimation of chemical substances by measure** Francis Sutton, 1871 Manual of Qualitative Chemical Analysis C. Remigius Fresenius, 1890 *Merck's Report* Theodore Weicker, 1902 *Journal of the Society of Chemical Industry* Society of Chemical Industry (Great Britain), 1895 **Calendar for the Year**, 1901 **Laboratory Manual of Elementary Chemistry** Jabez Montgomery, Roy Burnett Smith, 1904 The Chemical News and Journal of Physical Science

,1904

Unveiling the Magic of Words: A Report on "**Chemistry Practical Qualitative Analysis Sheet**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Chemistry Practical Qualitative Analysis Sheet**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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