

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

1. Calculate the number of moles of potassium chlorate, KClO_3 (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

Equation:	$2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$		
Before:	1.2	0	0
Change:	-1.2	+1.2	+1.8
After:	0	1.2	1.8

1.2 moles
 KClO_3 needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equation:	$\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$		
Before:	0.100	?	0
Change:	-0.100	-0.200 +	0.100 + 0.100
After:	0	?	0.100 0.100

0.200 moles
 HCl needed

3. Ethane, C_2H_6 reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	$2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$		
Before:	2.20	xs	0
Change:	-2.20	-7.70	+4.40 +6.60
After:	0	xs	4.40 6.60

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

Equation:	$\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$		
Before:	0.0040	XS	0
Change:	-0.0040	-0.0080	+0.0040 +0.0080
After:	0	XS	0.0040 0.0080

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

Chemistry Unit 3 Energy Study Answers

Annelies Wilder-Smith



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Table of Contents Chemistry Unit 3 Energy Study Answers

1. Understanding the eBook Chemistry Unit 3 Energy Study Answers
 - The Rise of Digital Reading Chemistry Unit 3 Energy Study Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 3 Energy Study Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Unit 3 Energy Study Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 3 Energy Study Answers

- Personalized Recommendations
 - Chemistry Unit 3 Energy Study Answers User Reviews and Ratings
 - Chemistry Unit 3 Energy Study Answers and Bestseller Lists
5. Accessing Chemistry Unit 3 Energy Study Answers Free and Paid eBooks
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 - Chemistry Unit 3 Energy Study Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Unit 3 Energy Study Answers
 - Highlighting and Note-Taking Chemistry Unit 3 Energy Study Answers
 - Interactive Elements Chemistry Unit 3 Energy Study Answers
 8. Staying Engaged with Chemistry Unit 3 Energy Study Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Unit 3 Energy Study Answers
 9. Balancing eBooks and Physical Books Chemistry Unit 3 Energy Study Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Unit 3 Energy Study Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Unit 3 Energy Study Answers
 - Setting Reading Goals Chemistry Unit 3 Energy Study Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Unit 3 Energy Study Answers

- Fact-Checking eBook Content of Chemistry Unit 3 Energy Study Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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