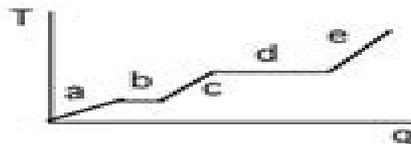


6. The highest vapor pressure is observed for which of the following liquid/temperature combinations?
- C_3H_8 at 270 K
 - C_2H_6 at 300 K
 - C_3H_{12} at 300 K
 - HOC_6H_5OH at 300 K
 - HOC_6H_5OH at 270 K



7. Which regions on the on the heating curve shown (Temperature versus heat, "q") corresponds to
- a pure gas increasing in temperature **e**
 - a pure liquid increasing in temperature **c**
 - a solid increasing in temperature **a**
 - a solid melting **b**
 - a liquid boiling **d**
8. On the above heating curve, which phases are present:
- In region "a" **solid**
 - In region "b" **solid/liquid**
 - In region "c" **liquid**
 - In region "d" **liquid/gas**
 - In region "e" **gas**
9. A solid is soft, and has a low melting point (below 300°C). Its solid, molten and aqueous forms are all nonconductors. Classify the solid.
10. Indicate the type of crystalline solid each of the following would form upon solidification. Identify the type of particles found at the lattice points and the types of attractive forces.

	Type of crystalline solid	Type of particles at lattice points	IMF
(a)	molecular solid	atoms	van der Waals
(b)	molecular	molecules	dipole-dipole
(c)	molecular	molecules	LDF
graphite	covalent network	atoms	covalent bond

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Ralph Thomas Myers



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