

# CHEM 121: Final Exam Study Guide

## Chapter 1

- Know the Scientific method
- Know the definitions for hypothesis, scientific law, and scientific theory.

## Chapter 2

- Length, mass, weight, volume
  - Know  $1 \text{ cm}^3 = 1 \text{ mL}$  and  $1 \text{ dm}^3 = 1 \text{ L}$
- Significant figures and Scientific notation
  - Rounding, in addition or subtraction, in multiplication or division
  - Keep as many sig figs until final answer
- Scientific notation
- Metric system
  - Know the metric prefixes: kilo (k), deci (d), centi (c), milli (m), and micro ( $\mu$ )
  - Be able to perform metric-metric conversions using these prefixes.
- Use the metric-English conversions provided ( $1 \text{ in.} = 2.54 \text{ cm}$ ;  $1 \text{ lb} = 453.6 \text{ g}$ ;  $1 \text{ qt} = 946 \text{ mL}$ )
- Volume by calculation
  - $V_{\text{rectangular solid}} = \text{length} \times \text{width} \times \text{thickness}$
- Volume by displacement
- Density:  $d = \frac{m}{V}$ 
  - Calculate density, mass, or volume
  - Identify what items sink or float given densities of liquids and solids.
- Temperature
  - Know the formulas for converting "F-to-"C or "C-to-"F and K-to-"C or "C-to-K
- Percentage: ratio of parts per 100 parts
  - Given amount of part and whole, calculate %
  - Use a given % to solve for part or whole

## Chapter 3

- Know that matter is studied at the macroscopic, microscopic, particulate (molecular) levels
- Physical states of matter
  - Determine physical state of substances (solids, liquids, gases) given descriptions of volume, shape, particles moving, etc.
- Classify properties and changes as physical or chemical
  - Know terms for changes of state:
    - Melting, freezing, vaporizing, condensation, sublimation, deposition
- Classification of matter
  - Given examples, determine which are elements, compounds, or mixtures
  - Given molecular-level images, determine which are elements, compounds, or mixtures and solids, liquids, or gases
  - Distinguish between homogeneous and heterogeneous mixtures
- Chemical reaction:
  - reactants: starting materials
  - products: substances produced in reaction

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**Southwestern Indian Polytechnic  
Institute**



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