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 $\text{GaBr} + \text{NaBr} + \text{SnO}_4 \rightarrow 4\text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{SO}_2 + 3\text{H}_2\text{O}$
 $5\text{NF}_3 + \text{SnF}_2 \rightarrow \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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 $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $2\text{ZnCl}_2 + 2\text{Cu} \rightarrow$ how many moles of $\text{Ce} + \text{ZnCl}_2 + \text{ZnCl}_2$ will be produced from 23.0 g of $\text{Ce} + \text{Zn}$ assuming $\text{Ce} + \text{CuCl}_2 + \text{CuCl}_2$ is available in excess

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