

## Balancing Equations Worksheet

- 1)  $\underline{\quad} \text{Na}_3\text{PO}_4 + \underline{\quad} \text{KOH} \rightarrow \underline{\quad} \text{NaOH} + \underline{\quad} \text{K}_3\text{PO}_4$
- 2)  $\underline{\quad} \text{MgF}_2 + \underline{\quad} \text{Li}_2\text{CO}_3 \rightarrow \underline{\quad} \text{MgCO}_3 + \underline{\quad} \text{LiF}$
- 3)  $\underline{\quad} \text{P}_4 + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{P}_2\text{O}_3$
- 4)  $\underline{\quad} \text{RbNO}_3 + \underline{\quad} \text{BeF}_2 \rightarrow \underline{\quad} \text{Be}(\text{NO}_3)_2 + \underline{\quad} \text{RbF}$
- 5)  $\underline{\quad} \text{AgNO}_3 + \underline{\quad} \text{Cu} \rightarrow \underline{\quad} \text{Cu}(\text{NO}_3)_2 + \underline{\quad} \text{Ag}$
- 6)  $\underline{\quad} \text{CF}_4 + \underline{\quad} \text{Br}_2 \rightarrow \underline{\quad} \text{CBr}_4 + \underline{\quad} \text{F}_2$
- 7)  $\underline{\quad} \text{HCN} + \underline{\quad} \text{CuSO}_4 \rightarrow \underline{\quad} \text{H}_2\text{SO}_4 + \underline{\quad} \text{Cu}(\text{CN})_2$
- 8)  $\underline{\quad} \text{GaF}_3 + \underline{\quad} \text{Cs} \rightarrow \underline{\quad} \text{CsF} + \underline{\quad} \text{Ga}$
- 9)  $\underline{\quad} \text{BaS} + \underline{\quad} \text{PtF}_2 \rightarrow \underline{\quad} \text{BaF}_2 + \underline{\quad} \text{PtS}$
- 10)  $\underline{\quad} \text{N}_2 + \underline{\quad} \text{H}_2 \rightarrow \underline{\quad} \text{NH}_3$
- 11)  $\underline{\quad} \text{NaF} + \underline{\quad} \text{Br}_2 \rightarrow \underline{\quad} \text{NaBr} + \underline{\quad} \text{F}_2$
- 12)  $\underline{\quad} \text{Pb}(\text{OH})_2 + \underline{\quad} \text{HCl} \rightarrow \underline{\quad} \text{H}_2\text{O} + \underline{\quad} \text{PbCl}_2$
- 13)  $\underline{\quad} \text{AlBr}_3 + \underline{\quad} \text{K}_2\text{SO}_4 \rightarrow \underline{\quad} \text{KBr} + \underline{\quad} \text{Al}_2(\text{SO}_4)_3$
- 14)  $\underline{\quad} \text{CH}_4 + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{CO}_2 + \underline{\quad} \text{H}_2\text{O}$
- 15)  $\underline{\quad} \text{Na}_3\text{PO}_4 + \underline{\quad} \text{CaCl}_2 \rightarrow \underline{\quad} \text{NaCl} + \underline{\quad} \text{Ca}_3(\text{PO}_4)_2$
- 16)  $\underline{\quad} \text{K} + \underline{\quad} \text{Cl}_2 \rightarrow \underline{\quad} \text{KCl}$
- 17)  $\underline{\quad} \text{Al} + \underline{\quad} \text{HCl} \rightarrow \underline{\quad} \text{H}_2 + \underline{\quad} \text{AlCl}_3$
- 18)  $\underline{\quad} \text{N}_2 + \underline{\quad} \text{F}_2 \rightarrow \underline{\quad} \text{NF}_3$
- 19)  $\underline{\quad} \text{SO}_2 + \underline{\quad} \text{Li}_2\text{Se} \rightarrow \underline{\quad} \text{SSe}_2 + \underline{\quad} \text{Li}_2\text{O}$
- 20)  $\underline{\quad} \text{NH}_3 + \underline{\quad} \text{H}_2\text{SO}_4 \rightarrow \underline{\quad} (\text{NH}_4)_2\text{SO}_4$

# Balancing Equation Practice Packet Answer Sheet

**David Kirk**



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web dec 8 2019 1 start by drawing a circle in the middle of your paper 2 then draw the snout of the unicorn start at bottom of the circle and curve the snout around to above the circle 3 then draw the front part of the unicorn s neck so that it is slanted 4 then draw the top of the unicorn s head

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