

AP Chemistry Test (Chapter 13)

Multiple Choice (30%)

- 1) Which one best describes the K_c for this reaction?



- A) $K_c = \frac{[A]^3[D]^2}{[E]^4[D]^2}$ B) $K_c = \frac{[B]^3[E]^4[G]^2}{[A]^3[D]^2[C]^5}$
C) $K_c = \frac{[E]^4[B]^3}{[A]^3[D]^2}$ D) $K_c = \frac{[A]^3[D]^2[C]^5}{[B]^3[E]^4[G]^2}$
E) None of these is correct

- 2) Please consider the gas phase reaction:



Which one would decrease the yield of H_2 ?

- A) High T & high P B) Low T & low P
C) High T & low P D) Low T & high P
3) $AlBr_3$ dissolves in water, according to this mechanism:



What effect would the addition of Na_2PO_4 have on this equilibrium?

- A) Q becomes larger & the reaction creates more products.
B) Q becomes larger & the reaction creates more reactants.
C) Q becomes smaller & the reaction creates more products.
D) Q becomes smaller & the reaction creates more reactants.
E) There is no net effect on the equilibrium position.
4) Based upon the following information, which one is true about the reaction mixture?



$$P (NH_3) = 2.19 \text{ atm} \quad P (N_2) = 4.33 \text{ atm} \quad P (H_2) = 0.72 \text{ atm}$$

- A) The product concentrations are too high & the reaction will shift left to establish equilibrium.
B) The product concentrations are too high & the reaction will shift right to establish equilibrium.
C) The reactant concentrations are too high & the reaction will shift right to establish equilibrium.
D) The reactant concentrations are too high & the reaction will shift left to establish equilibrium.

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S Ben Porath



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