

**AP<sup>®</sup> CHEMISTRY**  
**2011 SCORING GUIDELINES (Form B)**

**Question 1**  
**(10 points)**

Answer the following questions about the solubility and reactions of the ionic compounds  $M(OH)_2$  and  $MCO_3$ , where  $M$  represents an unidentified metal.

(a) Identify the charge of the  $M$  ion in the ionic compounds above.

2+	1 point is earned for the correct charge.
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(b) At 25°C, a saturated solution of  $M(OH)_2$  has a pH of 9.15.

(i) Calculate the molar concentration of  $OH^-(aq)$  in the saturated solution.

$pOH = 14 - pH$ $pOH = 14 - 9.15 = 4.85$ $[OH^-] = 10^{-4.85} = 1.4 \times 10^{-5} M$	1 point is earned for the correct concentration.
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(ii) Write the solubility-product constant expression for  $M(OH)_2$ .

$K_{sp} = [M^{2+}][OH^-]^2$	1 point is earned for the correct expression.
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(iii) Calculate the value of the solubility-product constant,  $K_{sp}$ , for  $M(OH)_2$  at 25°C.

$[M^{2+}] = \frac{1}{2} [OH^-] = \frac{1}{2} (1.4 \times 10^{-5} M) = 7.0 \times 10^{-6} M$ $K_{sp} = [M^{2+}][OH^-]^2 = (7.0 \times 10^{-6})(1.4 \times 10^{-5})^2$ $= 1.4 \times 10^{-15}$	1 point is earned for the correct relationship between $[M^{2+}]$ and $[OH^-]$ .  1 point is earned for the correct value.
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(c) For the metal carbonate,  $MCO_3$ , the value of the solubility-product constant,  $K_{sp}$ , is  $7.4 \times 10^{-14}$  at 25°C. On the basis of this information and your results in part (b), which compound,  $M(OH)_2$  or  $MCO_3$ , has the greater molar solubility in water at 25°C? Justify your answer with a calculation.

For $M(OH)_2$ : $[M^{2+}]$ and molar solubility = $7.0 \times 10^{-6} M$ For $MCO_3$ : $K_{sp} = 7.4 \times 10^{-14} = [M^{2+}][CO_3^{2-}]$ $[M^{2+}]$ and molar solubility = $2.7 \times 10^{-7} M$ Because $7.0 \times 10^{-6} M > 2.7 \times 10^{-7} M$ , $M(OH)_2$ has the greater molar solubility.	1 point is earned for the molar solubility of $MCO_3$ .  1 point is earned for an answer consistent with the calculated molar solubility.
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# Ap Chemistry 2011 Response Scoring Guidelines

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