

Name _____

SHORT ANSWER. Write the word or phrase that best completes each statement or answers the question.

- 1) What is the solvent in an aqueous solution?

ESSAY. Write your answer in the space provided or on a separate sheet of paper.

- 2) Explain why water is a good solvent for ionic substances.

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- 3) Of the species below, only _____ is
- not
- an electrolyte.

A) HCl B) Ar C) NaCl D) KOH E) Rb₂SO₄

- 4) What is the molarity of an aqueous solution containing 75.3 g of glucose (C
- ₆
- H
- ₁₂
- O
- ₆
-) in 35.5 mL of solution?

A) 0.197 B) 2.12 C) 3.52 D) 1.85 E) 11.8

- 5) The concentration of species in 500 mL of a 2.104 M solution of sodium sulfate is _____ M sodium ion and _____ M sulfate ion.

A) 4.208, 2.104 B) 1.052, 1.052 C) 2.104, 1.052 D) 2.104, 4.208 E) 2.104, 2.104

- 6) What is the molarity of a solution prepared by diluting 43.72 mL of 1.005 M aqueous K
- ₂
- Cr
- ₂
- O
- ₇
- to 500 mL?

A) 0.879 B) 87.9 C) 0.0879 D) 0.0115 E) 0.0218

- 7) When aqueous solutions of _____ are mixed, a precipitate forms.

A) NaI and KBr
B) Li₂CO₃ and CsI
C) K₂SO₄ and CrCl₃
D) NiBr₂ and AgNO₃
E) KOH and Ba(NO₃)₂

- 8) What are the spectator ions in the reaction between KOH (aq) and HNO
- ₃
- (aq)?

A) H⁺ and NO₃⁻
B) OH⁻ only
C) K⁺ and NO₃⁻
D) K⁺ and H⁺
E) H⁺ and OH⁻**SHORT ANSWER.** Write the word or phrase that best completes each statement or answers the question.

- 9) What is the concentration of HCl in a 0.100 M solution of HCl? Why?

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