

Terms to Know:

Electrochemistry – the study of the interchange of chemical and electrical energy

Voltaic or Galvanic Cell – IS a battery but not a dry cell; *generates useful electrical energy*

Electrolytic Cell – *requires useful electrical energy to drive a thermodynamically unfavorable reaction*

OIL RIG – oxidation is loss, reduction is gain (of electrons)

Oxidation – the loss of electrons, increase in charge

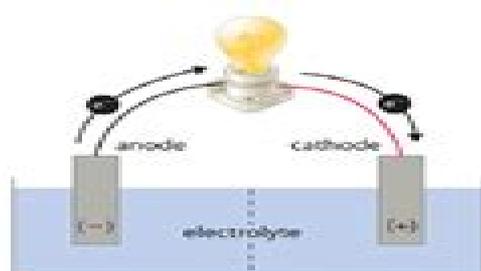
Reduction – the gain of electrons, reduction of charge

Oxidation number – the assigned charge on an atom

Electrochemistry Involves TWO MAIN TYPES Of Electrochemical Cells:

1. **Galvanic (voltaic) cells** – which are thermodynamically favorable chemical reactions (battery)
2. **Electrolytic cells** – which are thermodynamically unfavorable and require external e^- source (a direct current or DC power source)

BOTH of these fit into the category entitled **Electrochemical cells**



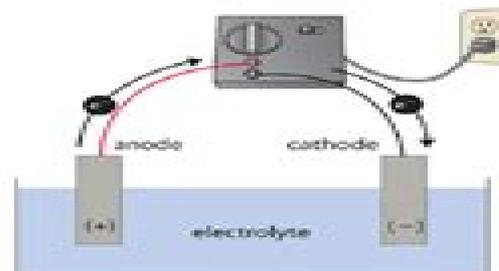
GALVANIC CELL

Energy released by spontaneous redox reaction is converted to electrical energy.

Oxidation half-reaction:
 $Y \rightarrow Y^+ + e^-$

Reduction half-reaction:
 $Z + e^- \rightarrow Z^-$

Overall cell reaction:
 $Y + Z \rightarrow Y^+ + Z^- \quad \Delta G < 0$



ELECTROLYTIC CELL

Electrical energy is used to drive nonspontaneous redox reaction.

Oxidation half-reaction:
 $Z^- \rightarrow Z + e^-$

Reduction half-reaction:
 $Y^+ + e^- \rightarrow Y$

Overall cell reaction:
 $Y^+ + Z^- \rightarrow Y + Z \quad \Delta G > 0$

Ap Chemistry Electrochemistry Answers

Mary Millhollon, Richard H. Langley



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Physical Chemistry (Solutions, Photochemistry, Electrochemistry) Dr. S. Ramakrishna, Dr. P.

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