

Revised August 2012

AP LAB 5c: REDOX Titration Simulations

Manganate(VII)/Fe²⁺ titration

1. Given that aqueous manganate(VII) (permanganate) ions will be converted to Mn²⁺ ions in acid solution, write a half equation to summarize this process.
2. Write a half equation to summarize the conversion of Fe²⁺ to Fe³⁺.
3. Combine the equations in #1 and #2 above to form a complete REDOX equation, and clearly identify the oxidizing agent and the reducing agent.
4. The REDOX titration computer simulation to accompany this LAB is located at www.chem.iastate.edu/group/Greenbowe/sections/projectfolder/flashfiles/redoxNew/redox.html

Go to the computer simulation and do the following:

- (a) Select the reaction: Choose KMnO₄. The buret will automatically fill with the selected oxidizing agent. The computer will automatically assign a molarity for the oxidizing agent.
 - (b) Use the slider to add a few mL of the oxidizing agent to the Fe²⁺ solution. Repeat this process until a slight pink "flash" of color is observed and then disappears. This is the signal that you are near the endpoint of the titration.
 - (c) Titrate to the end point. Getting the exact end point, requires great care (adding dropwise) and will be reached when there is a permanent pale pink color in the flask. Avoid adding too much oxidizing agent (dark purple) and having to start over.
5. Fill in the table below.

Molarity of oxidizing agent	
Volume of oxidizing agent added	
Volume of Fe ²⁺ solution	

6. Calculate the molarity of the Fe²⁺, enter this number (to three sig. figs) and click-on "OK". If you have done the titration accurately and correctly completed the calculation then you will have the correct answer. If NOT, then repeat the titration until the "Correct" message is seen.

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Rachel Sandford



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Table of Contents Ap Lab 5c Redox Titration Simulations Adrian Dingles

1. Understanding the eBook Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - The Rise of Digital Reading Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Personalized Recommendations
 - Ap Lab 5c Redox Titration Simulations Adrian Dingles User Reviews and Ratings
 - Ap Lab 5c Redox Titration Simulations Adrian Dingles and Bestseller Lists
5. Accessing Ap Lab 5c Redox Titration Simulations Adrian Dingles Free and Paid eBooks
 - Ap Lab 5c Redox Titration Simulations Adrian Dingles Public Domain eBooks
 - Ap Lab 5c Redox Titration Simulations Adrian Dingles eBook Subscription Services
 - Ap Lab 5c Redox Titration Simulations Adrian Dingles Budget-Friendly Options

6. Navigating Ap Lab 5c Redox Titration Simulations Adrian Dingles eBook Formats
 - ePub, PDF, MOBI, and More
 - Ap Lab 5c Redox Titration Simulations Adrian Dingles Compatibility with Devices
 - Ap Lab 5c Redox Titration Simulations Adrian Dingles Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Highlighting and Note-Taking Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Interactive Elements Ap Lab 5c Redox Titration Simulations Adrian Dingles
8. Staying Engaged with Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ap Lab 5c Redox Titration Simulations Adrian Dingles
9. Balancing eBooks and Physical Books Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ap Lab 5c Redox Titration Simulations Adrian Dingles
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Setting Reading Goals Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Fact-Checking eBook Content of Ap Lab 5c Redox Titration Simulations Adrian Dingles
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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