

1.4 Gli elementi

Gli elementi sono sostanze, come l'idrogeno e l'ossigeno, composte da un solo tipo di atomo. L'atomo è la particella più piccola di un elemento che conserva le proprietà caratteristiche dell'elemento stesso.

1.5 I composti

Una sostanza composta da elementi diversi tenuti insieme da un legame chimico (acqua, sale, ...) viene chiamata composto chimico.

Esistono due principali differenze tra una miscela e un composto puro: un composto presenta caratteristiche radicalmente differenti da quelle degli elementi che lo compongono e ha una definita composizione percentuale (in massa) degli elementi di cui è formato.

Alcuni composti (NaCl) sono formati da ioni, atomi o gruppi di atomi dotati di carica elettrica. Altri (zucchero, acqua) sono invece formati da molecole, che conservano le caratteristiche chimiche e la composizione del composto.

La composizione è rappresentata dalla relativa formula chimica.

1.6 Le proprietà fisiche

Le proprietà fisiche sono quelle proprietà che possono essere osservate e misurate senza variare la composizione della sostanza.

La temperatura spesso influenza i valori numerici delle proprietà di una sostanza. Ad es. H_2O varia la sua densità.

Le proprietà principalmente considerate in chimica sono:

- Colore
- Stato di aggregazione
- Punto di fusione
- Punto di ebollizione
- Densità
- Solubilità
- Conduttività elettrica
- Reattività
- Durezza
- Viscosità

Proprietà estensive ed intensive Le proprietà estensive dipendono dalla quantità di sostanza, mentre quelle intensive no. La densità, ad esempio, è una proprietà intensiva.

Appunti Di Chimica Generale Unife

RM Cervero



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