

Limitations : (i) During anodic protection method, the corrosion does take place, but at a very slow rate. In other words, corrosion rate is not zero. (ii) This method can be applied only to those metals/alloys which passivate.

Applications : Stainless steel containers used for transporting corrosive chemicals (e.g., concentrated acids) are protected from corrosion by passivating them by maintaining them at proper E_f and I_p .

Short Answer Questions

1. Define corrosion.

(Bombay, May 93, Nov. 97)

Ans. Any process of deterioration and consequent loss of solid metallic materials through an unwanted chemical or electrochemical attack by its environment, starting at its surface, is called corrosion.

2. What is meant by rusting of iron ?

Ans. The atmospheric corrosion of iron and steel, leading to the formation of a layer of reddish scale and powder of oxide (Fe_2O_3) on the surface.

3. What is rust ?

Ans. Hydrated ferric oxide (Fe_2O_3) with $Fe(OH)_3$ - Brown rust ; mixed ferrous and ferric oxides ($Fe_3O_4 = FeO \cdot Fe_2O_3$) - Black rust.

4. What is dry corrosion ?

Ans. Corrosion caused by direct chemical action of atmospheric gases such as oxygen, halogen, sulphur dioxide, hydrogen sulphide with metal surfaces in immediate proximity.

5. Formation of which types of metal oxide film cause rapid and continuous corrosion.

Ans. Volatile oxide film and porous oxide film.

6. Formation of which types of metal oxide film prevents corrosion.

Ans. Fine-grained tightly adhering, impervious oxide film ; and highly unstable oxide film.

7. What is Pilling-Bedworth rule ?

Ans. An oxide film is protective (or non-porous) if the volume of the oxide is at-least as great as the volume of the metal from which it is formed. On the other hand, if the volume of the oxide film is less than the volume of the metal, then the oxide film is porous and non-protective.

8. What is wet corrosion ?

Ans. A corrosion which occurs when a conducting liquid is in contact with : (i) metal or (ii) when two dissimilar metals or alloys are either immersed or dipped partially in a solution.

9. What is electrochemical corrosion ?

Ans. This type of corrosion occurs due to the existence of separate 'anodic' and 'cathodic' areas between which current flows through the conducting solution. At anodic area oxidation occurs and anodic part of metal is destroyed.

10. What is galvanic corrosion ?

Ans. When two dissimilar metals are electrically connected and exposed to an electrolyte, the metal higher in electrochemical series undergoes corrosion.

11. The rate of metallic corrosion increases with increase in temperature. Give reason.

(Bangalore, Aug. 95)

Ans. With increases of temperature of the environment, the rate of reaction as well as rate of diffusion increase, thereby corrosion rate increases.

12. Iron corrodes faster than aluminium, even though iron is placed below aluminium in the electrical series. Why ?

(Bangalore, Jan. 93, Aug. 95)

Ans. This is because aluminium forms a non-porous, very thin, tightly adhering protective oxide film (Al_2O_3) on its surface and this film does not permit corrosion to occur.

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Table of Contents Jain And Jain 15th Edition

1. Understanding the eBook Jain And Jain 15th Edition
 - The Rise of Digital Reading Jain And Jain 15th Edition
 - Advantages of eBooks Over Traditional Books
2. Identifying Jain And Jain 15th Edition
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Jain And Jain 15th Edition
 - User-Friendly Interface

4. Exploring eBook Recommendations from Jain And Jain 15th Edition
 - Personalized Recommendations
 - Jain And Jain 15th Edition User Reviews and Ratings
 - Jain And Jain 15th Edition and Bestseller Lists
5. Accessing Jain And Jain 15th Edition Free and Paid eBooks
 - Jain And Jain 15th Edition Public Domain eBooks
 - Jain And Jain 15th Edition eBook Subscription Services
 - Jain And Jain 15th Edition Budget-Friendly Options
6. Navigating Jain And Jain 15th Edition eBook Formats
 - ePub, PDF, MOBI, and More
 - Jain And Jain 15th Edition Compatibility with Devices
 - Jain And Jain 15th Edition Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Jain And Jain 15th Edition
 - Highlighting and Note-Taking Jain And Jain 15th Edition
 - Interactive Elements Jain And Jain 15th Edition
8. Staying Engaged with Jain And Jain 15th Edition
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Jain And Jain 15th Edition
9. Balancing eBooks and Physical Books Jain And Jain 15th Edition
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Jain And Jain 15th Edition
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Jain And Jain 15th Edition
 - Setting Reading Goals Jain And Jain 15th Edition
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Jain And Jain 15th Edition
 - Fact-Checking eBook Content of Jain And Jain 15th Edition
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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