

AP CHEMISTRY CHAPTER 6 PRACTICE TEST

- For which chemical equation will the heat of reaction  $\Delta H^\circ$ , correspond to the molar heat of formation of liquid water?
  - $\text{H}_{2(g)} + 1/2\text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$
  - $\text{H}_{(g)} + \text{H}_{(g)} + \text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)}$
  - $2\text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2\text{H}_2\text{O}_{(l)}$
  - $\text{H}^+_{(aq)} + \text{OH}^-_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)}$
- For the process  $\text{H}_{2(g)} + 1/2\text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$ ,  $\Delta H = -247$  kJ. The numeric value of -242 kJ represents heat energy that is
  - absorbed by the reaction
  - required to break the H-H and O=O bond
  - stored as potential energy
  - released by the reaction
- An 11.0 g sample of propane is placed in a bomb calorimeter with excess oxygen and is ignited. The initial temperature of the calorimeter is 25.00°C and its total heat capacity is 97.1 kJ/°C. The reaction raises the temperature of the calorimeter is 30.75°C. The approximate molar heat of combustion for propane is
  - 50.6 kJ/mol
  - 558 kJ/mol
  - $-2.24 \times 10^3$  kJ/mol
  - $-1.20 \times 10^4$  kJ/mol
- What is the change in enthalpy at 25°C and 1 atm for the production of 3.00 mol  $\text{SnO}_{(s)}$ ?
 
$$\text{Sn}_{(s)} + \text{SnO}_{2(s)} \rightarrow 2\text{SnO}_{(s)} \quad \Delta H^\circ = 16.2 \text{ kJ}$$
  - 24.3 kJ
  - 16.2 kJ
  - 5.40 kJ
  - 16.2 kJ
  - 24.3 kJ
- How much heat is liberated at constant pressure if 0.896 g of calcium carbonate reacts with 92.3 mL of 0.584 M hydrochloric acid?
 
$$\text{CaCO}_{3(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{CaCl}_{2(aq)} + \text{H}_2\text{O}_{(l)} + \text{CO}_{2(g)} \quad \Delta H^\circ = -15.2 \text{ kJ}$$
  - 0.410 kJ
  - 8.88 kJ
  - 13.6 kJ
  - 0.546 kJ
  - 0.136 kJ

# Ap Chemistry Chapter 6 Test

**John T. Moore, Richard H. Langley**



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## **Table of Contents Ap Chemistry Chapter 6 Test**

1. Understanding the eBook Ap Chemistry Chapter 6 Test
  - The Rise of Digital Reading Ap Chemistry Chapter 6 Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Chapter 6 Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ap Chemistry Chapter 6 Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Chapter 6 Test
  - Personalized Recommendations
  - Ap Chemistry Chapter 6 Test User Reviews and Ratings
  - Ap Chemistry Chapter 6 Test and Bestseller Lists
5. Accessing Ap Chemistry Chapter 6 Test Free and Paid eBooks
  - Ap Chemistry Chapter 6 Test Public Domain eBooks
  - Ap Chemistry Chapter 6 Test eBook Subscription Services
  - Ap Chemistry Chapter 6 Test Budget-Friendly Options
6. Navigating Ap Chemistry Chapter 6 Test eBook Formats
  - ePub, PDF, MOBI, and More
  - Ap Chemistry Chapter 6 Test Compatibility with Devices
  - Ap Chemistry Chapter 6 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ap Chemistry Chapter 6 Test
  - Highlighting and Note-Taking Ap Chemistry Chapter 6 Test
  - Interactive Elements Ap Chemistry Chapter 6 Test
8. Staying Engaged with Ap Chemistry Chapter 6 Test

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ap Chemistry Chapter 6 Test
9. Balancing eBooks and Physical Books Ap Chemistry Chapter 6 Test
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ap Chemistry Chapter 6 Test
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ap Chemistry Chapter 6 Test
- Setting Reading Goals Ap Chemistry Chapter 6 Test
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chemistry Chapter 6 Test
- Fact-Checking eBook Content of Ap Chemistry Chapter 6 Test
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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