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web 33mol O_2 6mol Br_2 11mol O_2 18mol Br_2 identify the limiting reactant s and excess reactant s the limiting reactant is O_2 since it would yield the least amount of product 18 mol Br_2 the excess reactant is $\text{C}_2\text{H}_3\text{Br}_3$ since its complete reaction would have yielded up to 54 mol Br_2 think about your result

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web step 4 the reactant that produces a smaller amount of product is the limiting reagent mg mg produces less MgO MgO than does O_2 0.2398 g MgO vs 25.2 g MgO therefore mg is the limiting reagent in this reaction step 5 the reactant that produces a larger amount of product is the excess reagent

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web please save your changes before editing any questions 5 minutes 1 pt viewing model 3 if 4 moles of hydrogen react with 2 moles of oxygen which of these is the excess reactant is leftover at the end and how many moles of it are leftover hydrogen 2 moles leftover hydrogen 1 mole leftover

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web if you start with 14 82 g of ca oh 2 c a o h 2 and 16 35 g of h2so4 h 2 s o 4 a determine the limiting reagent b determine the number of moles of h2o h 2 o produced c determine the number of grams of caso4 c a s o 4 produced d determine the number of grams of excess reagent left 1 make sure the equation is balanced

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