

# UNIT $\Rightarrow$ 1

## ATOMIC & MOLECULAR STRUCTURE

- ⊙ s-orbital  $\rightarrow$   $\oplus$   $\ominus$  Spherical shape
- ⊙ p-orbital  $\rightarrow$   $\infty$  dumbbell shape
- ⊙ वै orbitals तनेगा जिसका atomic energy same होगा
- ⊙ जब molecules बनता है तब e की जगह नहीं होता है केवल orbital overlap होता है।

$$\text{No. of atomic orbitals} = \text{No. of molecular orbitals}$$

### # Molecular orbital theory -

In the molecular orbital theory the valence electrons are considered to be associated with all the nuclei in the molecule.

### Postulates -

- $\Rightarrow$  when nuclei of two atoms comes close to each other, their atomic orbitals interact to form molecular orbital in

# B Tech Semester 1st 2nd Engineering Chemistry Ch 101

**M Tight**



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