

Soluzioni

CAPITOLO 16

VERIFICA LE TUE CONOSCENZE

LE EQUAZIONI CHIMICHE

- 1** I simboli convenzionali utilizzati nella rappresentazione delle equazioni chimiche sono: le formule chimiche dei reagenti e dei prodotti, la freccia di reazione singola (\rightarrow) o doppia (\rightleftharpoons), i simboli che indicano la produzione di un precipitato (\downarrow), lo sviluppo di un gas (\uparrow), l'aggiunta di calore (Δ). Infine i simboli che indicano lo stato fisico dei reagenti e dei prodotti.
- 2** a. $N_{2(g)} + 3 H_{2(g)} \rightarrow 2 NH_{3(g)}$
- 3** b. Fosforo
- 4** d. $Al + 2 HNO_3 \rightarrow Al(NO_3)_3 + H_2$
- 5** c. 1,1 \rightarrow 1,3

LE REAZIONI DI SINTESI E ANALISI, COMBINAZIONE E DECOMPOSIZIONE

- 6** 1. Metallo + ossigeno \rightarrow ossido basico;
2. non metallo + ossigeno \rightarrow ossido acido;
3. metallo + idrogeno \rightarrow idruro ionico;
4. non metallo + idrogeno \rightarrow idruro covalente oppure idracido;
5. metallo + non metallo \rightarrow sali di idracidi.
- 7**
a. Ossoacido: ossido acido + acqua;
b. carbonato: ossido basico + CO_2 .
- 8** a. un ossido acido.
- 9** In acqua, forma un acido: c. SO_3 .
- 10** Reazione di: d. decomposizione.
- 11** Prodotto della reazione tra litio metallico e cloro gassoso: b. LiCl.

REAZIONI DI SCAMBIO SEMPLICE

- 12** 1. Metallo + acqua \rightarrow idrossido + idrogeno;
2. metallo + acido \rightarrow sale + idrogeno;
3. metallo + sale di ossoacido \rightarrow sale di ossoacido + metallo.
- 13** Un prodotto della reazione è: b. idrogeno.
- 14** Non reagisce con l'acqua: c. Cu(II).
- 15** c. alluminio e acido carbonico.

LE REAZIONI DI DOPPIO SCAMBIO

- 16** 1. Sale (solfuro o cloruro) + acido \rightarrow sale + solfuro o cloruro di idrogeno (\uparrow);
2. sale (carbonato o solfito) + acido \rightarrow sale + diossido di carbonio o diossido di zolfo (\uparrow) + H_2O ;
3. sale + acido \rightarrow acido + sale (\downarrow);
4. sale + sale \rightarrow sale + sale (\downarrow);
5. sale + base \rightarrow sale + base (\downarrow);
6. sale di ammonio + base \rightarrow sale + NH_3 + H_2O .
- 17** Composto più solubile in acqua:
a. $BaSO_4 / K_2S / Mg_3(PO_4)_2 / K_2S$
b. $AgCl / Fe(OH)_3 / Na_2Cr_2O_7 / Na_2Cr_2O_7$
c. $Fe(NO_3)_2 / PbCO_3 / Ag_2CrO_4 / Fe(NO_3)_2$
- 18** Il più solubile in acqua è: b. solfato rameico.
- 19** Nella reazione tra un acido e: b. solfito di sodio.

LE REAZIONI ACIDO-BASE

- 20** 1. Ossido basico + acido \rightarrow sale + H_2O ;
2. idrossido + ossido acido \rightarrow sale + H_2O ;
3. idrossido + acido \rightarrow sale + H_2O .
- 21** a. $2 NaOH + H_2S \rightarrow Na_2S + 2 H_2O$
- 22** Sei molecole di acqua e: b. fosfato di calcio.
- 23** È necessario far reagire: a. HCl e $Al(OH)_3$.

L'EQUAZIONE IONICA NETTA

- 24** Nella reazione tra un idrossido e un acido, la reazione avviene tra gli ioni H^+ rilasciati dall'acido e gli ioni OH^- rilasciati dalla base:
 $2 H^+_{(aq)} + 2 OH^-_{(aq)} \rightarrow 2 H_2O_{(l)}$
- 25**
a. $Cu^{2+}_{(aq)} + OH^-_{(aq)} \rightarrow Cu(OH)_{2(aq)}$
b. $2 Ag^+_{(aq)} + S^{2-}_{(aq)} \rightarrow Ag_2S_{(s)}$
c. $3 Mg^{2+}_{(aq)} + 2 PO_4^{3-}_{(aq)} \rightarrow Mg_3(PO_4)_{2(aq)}$
d. $Pb^{2+}_{(aq)} + SO_4^{2-}_{(aq)} \rightarrow PbSO_{4(aq)}$
- 26** Ioni spettatori: a. $Na^+_{(aq)}$ e $Cl^-_{(aq)}$
- 27** Il precipitato che si forma è: b. $FePO_4$.

LE QUANTITÀ DI REAGENTI E PRODOTTI

- 28** La stechiometria di una reazione chimica definisce: b. l'aspetto quantitativo.
- 29** Completamento errato il numero di a. ioni.

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L Manion



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