

PHYSICAL PHARMACY (LECTURE)

Rays-J Leradd, RPh, MSPhar

IMPORTANCE OF STUDYING THE PHENOMENON OF SOLUBILITY

Understanding the phenomenon of solubility helps the pharmacist to:

1. Select the best solvent for a drug or a mixture of drugs.
2. Overcome problems arising during preparation of pharmaceutical solutions.
3. Have information about the structure and intermolecular forces of the drug.
4. Many drugs are formulated as solutions or added as powder or solution forms to liquids.
5. Drugs with low aqueous solubility often present problems related to their formulation and bioavailability.

SOLUBILITY

- > **In a quantitative way:** it is the concentration of solute in a saturated solution at a certain temperature.
- > **In a qualitative way:** it is the spontaneous interaction of two or more substances (solute and solvent) to form a homogenous molecular dispersion.

DEFINITION OF TERMS

SOLUTION

- > is a mixture of two or more components that forms a homogenous mixture. The components are referred to the solute and/or solutes and the solvent and/or solvents.

SOLUTE

- > is the dissolved agent (less abundant part of the solution)

SOLVENT

- > is the component in which the solute is dissolved (more abundant part of the solution)

A SUPERSATURATED SOLUTION

- > contains more of the dissolved solute than it would normally contain in a saturated state at a definite temperature.

DEGREE OF SATURATION SOLUBILITY CURVE

Any solution can be saturated, unsaturated, or supersaturated by changing temperature.



THERMODYNAMIC SOLUBILITY OF DRUGS

- > The thermodynamic solubility of a drug in a solvent is the maximum amount of the most stable crystalline form that remains in solution (or quantity of the drug substance which can be completely dissolved) in a given volume of the solvent at a given temperature and pressure under equilibrium conditions.
- > The equilibrium involves a balance of the energy of three interactions against each other:
 - o Solvent with solvent
 - o Solute with solute
 - o Solvent and solute

Fig. 10. Solubility depends on the solution.



Fig. 11. Solubility of the solid depends on the solvent.



Fig. 12. The solid solubility depends on temperature and the solute in the solution.



Fig. 13. Process of solubilization.

SOLUBILITY PROCESS

A mechanistic perspective of solubilization process for organic solute in water involves the following steps:

1. Break up of solute-solute intermolecular bonds.
2. Break up of solvent-solvent intermolecular bonds.
3. Formation of cavity in solvent phase large enough to accommodate solute molecule.
4. Transfer of solute into the cavity of solvent phase.
5. Formation of solute-solvent intermolecular bonds.

ENTHALPY

- > The enthalpy change of solution refers to the overall amount of heat which is released or absorbed during the dissolving process (at constant pressure).
- > The enthalpy of solution can either be positive (endothermic reaction) or negative (exothermic reaction).



Physical Pharmacy Lecture Notes

**Anja Maier, Josef Oehmen, Pieter E.
Vermaas**

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