



NATIONAL
EXAM CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

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QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

QUESTION 3

QUESTION 3 starts on PAPER 2 page 2.

3.1 The boiling points in the table below were obtained during an investigation into the boiling points of pure 1-butanol. The table below lists the boiling points of pure 1-butanol and pure 1-butanol with 10% water.

Compound	Boiling point (°C)	Boiling point (°C)
A	117	117
B	117	117

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 In which column are compounds A and B in (1) and (2) lists of boiling points? (2)
- 3.1.3 Name the type of intermolecular force that water shows between molecules of pure compounds A and B due to the polar nature of their molecules. (2)
- 3.1.4 What type of intermolecular force is present in pure 1-butanol? (2)
- 3.1.5 What compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 Give a reason for the order in which the water is added. (2)

Compound B (10% of water in the table) has a smaller boiling point than pure compounds A and B but has a smaller vapour boiling point of 114 °C.

- 3.1.7 Explain why the boiling point of compound B is 114 °C (2) and why it is smaller than the boiling point of pure 1-butanol (2).

3.2 Additive D is added to each mixture of pure compounds A and B to lower its boiling point to the value in the table below.

10%	114	114
20%	114	114

- 3.2.1 In which column (1) or (2), will the water be added? (2)
- 3.2.2 Explain the order in (QUESTION 3.1) about the order in the intermolecular force tables. (2)

QUESTION 3



Physical Science Paper 2 Intermolecular Forces Investigation

JL Elias



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Proceedings of the Royal Society. Section A, Mathematical and Physical Science Royal Society (Great Britain),1908
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