

## LAWS OF THERMODYNAMICS

### 1 ZEROth LAW OF THERMODYNAMICS

*If the systems A and B are in thermal equilibrium separately and A and C are in thermal equilibrium separately, then B and C should be in thermal equilibrium.*

This law is the basis for temperature measurement. When two bodies at different temperatures are brought into contact, after some time, they attain a common temperature and are then said to exist in thermal equilibrium.

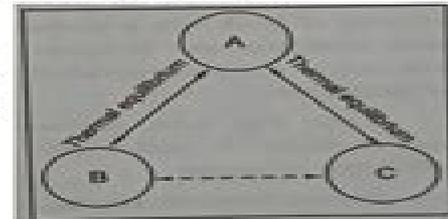


Fig. 1.1

### 2 FIRST LAW OF THERMODYNAMICS

The first law of thermodynamics evolving from **conservation of energy principle**, provides a sound basis for studying the relationships among the various forms of energy and energy interactions. Refer Fig. 1.2

A rock at some elevation possesses some potential energy, and part of this potential energy is converted into kinetic energy as the rock falls. Experimental data show that the decrease in potential energy ( $mg \Delta Z$ ) exactly equals to the increase in kinetic energy

$$\frac{m (v_f^2 - v_i^2)}{2}$$

when the air resistance is negligible;

The increase in the internal energy of a potato in an oven is equal to the amount of heat transferred to it.

**According to First Law of Thermodynamics**

*"Heat can be converted into work and work can be converted into heat."*

# Laws Of Thermodynamics In Mechanical Engineering

**Ensheng Dong**



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