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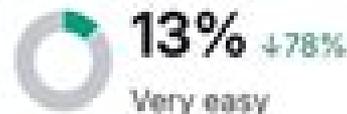
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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds $\text{Cl}(\text{aq})$

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 $13 \text{Ag}(\text{aq}) + 2\text{F}(\text{aq}) + 2\text{NH}_4(\text{aq}) + \text{Cr}_2\text{O}_7^{2-}(\text{aq}) \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7(\text{s}) + 2\text{NH}_4(\text{aq}) + 2\text{F}(\text{aq})$ because two $\text{NH}_4(\text{aq})$

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