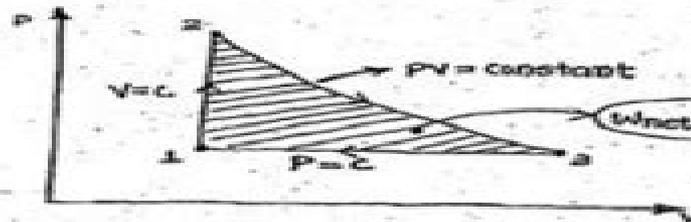


Solution (Assignment 1)

Given:

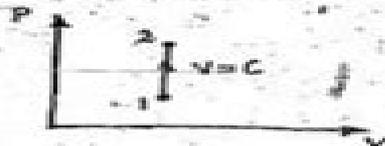
$T_1 = 350 \text{ K}$   
 $T_2 = 3T_1 = 3(350) = 1050 \text{ K}$   
 $T_3 = T_2$  (isothermal process)  
 $T_3 = 1050 \text{ K}$   
 $n = 1$



$V_2 = V_1$   
 $T_3 = T_2$

Net work done / kg of gas =  $w_{12} + w_{23} + w_{31}$

For process 1-2



constant volume process

$\frac{P_2}{P_1} = \frac{T_2}{T_1} = \frac{1050}{350}$

$\frac{P_2}{P_1} = 3$

$P_2 = 3P_1$

$w_{12} = 0$



isothermal process

$PV = \text{constant}$

$w_{23} = P_2 V_2 \ln \frac{V_3}{V_2}$   
 $= mRT_2 \ln \frac{V_3}{V_2}$   
 $= (1)(R)1050 \ln 3$   
 $= R(1050) \ln 3$   
 $= 1050 \times 1.0986 \times R$

$w_{23} = 1153.542909 R$



constant pressure process

$P = C$

$\frac{V_3}{V_1} = \frac{T_3}{T_1}$

$\frac{V_3}{V_1} = \frac{T_3}{T_1} = \frac{1050}{350}$

$\frac{V_3}{V_1} = 3$

$w_{31} = P(V_1 - V_3)$   
 $= P V_1 - P V_3$   
 $= mRT_1 - mRT_3$   
 $= RT_1 - RT_3$   
 $= R(T_1 - T_3)$   
 $= R(350 - 1050)$

$w_{31} = -700 R$

Net work done / kg of gas =  $w_{12} + w_{23} + w_{31}$

$= 0 + 1153.5429 R - 700 R$

$\frac{w_{\text{net}}}{\text{kg of gas}} = 453.5429 R \text{ Ans}$

# Thermodynamics Mechanical Engineering Notes

**RC Schank**



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