

## Assessment

**Chapter Test B****Chapter: Chemical Formulas and Chemical Compounds**

**PART I** In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. Changing a subscript in a correctly written chemical formula will
- change the electron configuration of that element.
  - change the charges on the other ions in the compound.
  - change the formula so that it no longer represents the same compound.
  - have no effect on the formula.
- \_\_\_\_\_ 2. Using the Stock system of nomenclature,  $\text{Cr}_2(\text{SO}_4)_3$  is named
- chromium(II) sulfate.
  - chromic sulfate.
  - dichromium trisulfate.
  - chromium(III) sulfate.
- \_\_\_\_\_ 3. In a polyatomic ion, the algebraic sum of the oxidation numbers of all atoms is equal to
- 0.
  - 10.
  - the number of atoms in the ion.
  - the charge on the ion.
- \_\_\_\_\_ 4. The first part of the name of a binary ionic compound is the
- cation.
  - polyatomic ion.
  - oxyanion.
  - anion.
- \_\_\_\_\_ 5. The empirical formula may not represent the actual composition of a(n)
- ionic compound.
  - crystal.
  - salt.
  - molecular compound.
- \_\_\_\_\_ 6. The formula for phosphoric acid is
- $\text{HPO}_4$
  - $\text{H}(\text{PO}_4)_2$
  - $\text{H}_2\text{PO}_3$
  - $\text{H}_3\text{PO}_4$

# Chapter 7 Chemical Formulas And Compounds Test B

**Ranjeet Shahi**



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## **Table of Contents Chapter 7 Chemical Formulas And Compounds Test B**

1. Understanding the eBook Chapter 7 Chemical Formulas And Compounds Test B
  - The Rise of Digital Reading Chapter 7 Chemical Formulas And Compounds Test B
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 7 Chemical Formulas And Compounds Test B
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 7 Chemical Formulas And Compounds Test B
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 7 Chemical Formulas And Compounds Test B
  - Personalized Recommendations
  - Chapter 7 Chemical Formulas And Compounds Test B User Reviews and Ratings
  - Chapter 7 Chemical Formulas And Compounds Test B and Bestseller Lists
5. Accessing Chapter 7 Chemical Formulas And Compounds Test B Free and Paid eBooks
  - Chapter 7 Chemical Formulas And Compounds Test B Public Domain eBooks
  - Chapter 7 Chemical Formulas And Compounds Test B eBook Subscription Services

- Chapter 7 Chemical Formulas And Compounds Test B Budget-Friendly Options
- 6. Navigating Chapter 7 Chemical Formulas And Compounds Test B eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 7 Chemical Formulas And Compounds Test B Compatibility with Devices
  - Chapter 7 Chemical Formulas And Compounds Test B Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 7 Chemical Formulas And Compounds Test B
  - Highlighting and Note-Taking Chapter 7 Chemical Formulas And Compounds Test B
  - Interactive Elements Chapter 7 Chemical Formulas And Compounds Test B
- 8. Staying Engaged with Chapter 7 Chemical Formulas And Compounds Test B
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 7 Chemical Formulas And Compounds Test B
- 9. Balancing eBooks and Physical Books Chapter 7 Chemical Formulas And Compounds Test B
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 7 Chemical Formulas And Compounds Test B
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 7 Chemical Formulas And Compounds Test B
  - Setting Reading Goals Chapter 7 Chemical Formulas And Compounds Test B
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 7 Chemical Formulas And Compounds Test B
  - Fact-Checking eBook Content of Chapter 7 Chemical Formulas And Compounds Test B
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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