

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
(a) formula unit. (c) crystal.
(b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
(a) molecules. (c) dipoles.
(b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
(a) higher in potential energy. (c) equal in potential energy.
(b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
(a) Compound A is not an ionic compound.
(b) It will be more difficult to break the bonds in compound A than those in compound B.
(c) Compound B has larger crystals than compound A.
(d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
(a) stronger than the attractive forces among formula units in ionic bonding.
(b) weaker than the attractive forces among formula units in ionic bonding.
(c) approximately equal to the attractive forces among formula units in ionic bonding.
(d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

- What type of energy best represents the strength of an ionic bond?

Lattice energy

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

Chemistry Chapter 6 Review Answers

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Independent and Weekly Review , 1914

Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

5 Steps to a 5 AP Chemistry John Moore, 2003-08-22 For the more than one million students taking the AP exams each year Boxed quotes offering advice from students who have aced the exams and from AP teachers and college professors Sample tests that closely simulate real exams Review material based on the contents of the most recent tests Icons highlighting important facts vocabulary and frequently asked questions Websites and links to valuable online test resources along with author e mail addresses for students with follow up questions Authors who are either AP course instructors or exam developers

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Aeronautical Engineering Review, 1953

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AP Chemistry Albert S. Tarendash, Frederick J. Rowe, 2002 This indispensable guide to chemistry helps students who wish to prepare for the AP Chemistry exam on their own Comprehensive and easy to understand this learning guide includes a full content review two full length practice tests with hundreds of practice questions and thorough answer explanations and proven test taking strategies

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