



CHEMISTRY

PERIODIC TABLE OF THE ELEMENTS

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
H	He																
Li	Be	B	C	N	O	F	Ne										
Na	Mg	Al	Si	P	S	Cl	Ar										
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra	Ac	Rf	Mn	Co	Ni	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe		

PERIODIC TRENDS

- Atomic Size**
 - Increases from top-left to bottom-right.
 - Increases from top to bottom.
 - Increases from left to right.
- Ionization Energy**
 - Increases from bottom-left to top-right.
 - Increases from left to right.
 - Increases from bottom to top.
- Electronegativity**
 - Increases from bottom-left to top-right.
 - Increases from left to right.
 - Increases from bottom to top.

ATOMIC STRUCTURE AND THE PERIODIC TABLE

- Atomic Structure**
 - Electrons are arranged in shells.
 - Electrons in the outer shell are valence electrons.
 - The number of valence electrons determines the chemical properties of an element.
- Periodic Trends**
 - Elements in the same group have similar chemical properties.
 - Elements in the same period have the same number of electron shells.

$1 - \frac{1}{n^2} = R_H \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$

where $R_H = 1.097 \times 10^7 \text{ m}^{-1}$, n_1 and n_2 are integers representing the initial and final energy levels of the electron.

IONIZATION ENERGY AND ELECTRONEGATIVITY



Ionization energy is the energy required to remove an electron from an atom. Electronegativity is the ability of an atom to attract electrons.

SHORTEST HISTORY OF ATOMS

- Democritus (470-370 BC) proposed the concept of atoms.
- Aristotle (384-322 BC) proposed the concept of elements.
- John Dalton (1766-1844) proposed the atomic theory.
- J.J. Thomson (1857-1940) discovered the electron.
- Ernest Rutherford (1871-1937) discovered the nucleus.
- Niels Bohr (1879-1962) proposed the Bohr model.
- Erwin Schrödinger (1927) proposed the quantum mechanical model.



THE NEW MECHANICAL MODEL

- The atom is a cloud of electrons.
- Electrons are found in orbitals.
- The probability of finding an electron in a certain region is given by the square of the wave function.

- The atom is a cloud of electrons.
- Electrons are found in orbitals.
- The probability of finding an electron in a certain region is given by the square of the wave function.

PERIODIC TABLE

- The periodic table is a table of elements.
- Elements are arranged in groups and periods.
- Groups are vertical columns.
- Periods are horizontal rows.
- Elements in the same group have similar chemical properties.
- Elements in the same period have the same number of electron shells.
- The periodic table is a key to understanding the properties of elements.

WORKING

- Work on the periodic table.
- Identify the groups and periods.
- Describe the properties of elements.
- Explain the trends in the periodic table.

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