

Mathematically

$$\frac{P^0 - P^s}{P^0} = x_2$$

Here  $P^0$  = Pressure of solvent in pure stage.

$P^s$  = Pressure of solution

$x_2$  = Mole fraction of solute.

### DERIVATION OF THE LAW

Pressure of solvent in solution

= Mole fraction of solvent  $\times$  pressure of solvent in pure stage

**If solute is non-volatile**

Total pressure of solution = Pressure due to solvent

= Mole fraction of solvent  $\times$  pressure of solvent in pure stage

So  $P_{\text{solvent}} = x_1 \cdot P_{\text{solution}}$  Here  $x_1$  is mole fraction of solvent.

$$\frac{P_{\text{solvent}}}{P_{\text{solution}}} = x_1$$

Subtract each term from 1 on both side.

$$1 - \frac{P_{\text{solvent}}}{P_{\text{solution}}} = 1 - x_1$$

$$\frac{P_{\text{solution}} - P_{\text{solvent}}}{P_{\text{solution}}} = x_2$$

### 2.11 IDEAL & NON-IDEAL SOLUTION

Sr.	Ideal solution	Non-Ideal solution
1.	A solution which obey Roult's law is called ideal solution.	A solution which do not obey Roult's law is called non-ideal solution.
2.	Interaction between components of mixture i.e.-between A&B is some as found in A-A & B-B molecule.	Interaction between components of mixture i.e.-between A&B is not some as found in A-A & B-B molecule.

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**Hussin A.Rothana**



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## **Table of Contents Chapter Wise Class 12 Chemistry Notes**

1. Understanding the eBook Chapter Wise Class 12 Chemistry Notes
  - The Rise of Digital Reading Chapter Wise Class 12 Chemistry Notes
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Wise Class 12 Chemistry Notes
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter Wise Class 12 Chemistry Notes
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Wise Class 12 Chemistry Notes
  - Personalized Recommendations
  - Chapter Wise Class 12 Chemistry Notes User Reviews and Ratings

- Chapter Wise Class 12 Chemistry Notes and Bestseller Lists
- 5. Accessing Chapter Wise Class 12 Chemistry Notes Free and Paid eBooks
  - Chapter Wise Class 12 Chemistry Notes Public Domain eBooks
  - Chapter Wise Class 12 Chemistry Notes eBook Subscription Services
  - Chapter Wise Class 12 Chemistry Notes Budget-Friendly Options
- 6. Navigating Chapter Wise Class 12 Chemistry Notes eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter Wise Class 12 Chemistry Notes Compatibility with Devices
  - Chapter Wise Class 12 Chemistry Notes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter Wise Class 12 Chemistry Notes
  - Highlighting and Note-Taking Chapter Wise Class 12 Chemistry Notes
  - Interactive Elements Chapter Wise Class 12 Chemistry Notes
- 8. Staying Engaged with Chapter Wise Class 12 Chemistry Notes
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter Wise Class 12 Chemistry Notes
- 9. Balancing eBooks and Physical Books Chapter Wise Class 12 Chemistry Notes
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter Wise Class 12 Chemistry Notes
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Wise Class 12 Chemistry Notes
  - Setting Reading Goals Chapter Wise Class 12 Chemistry Notes
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Wise Class 12 Chemistry Notes
  - Fact-Checking eBook Content of Chapter Wise Class 12 Chemistry Notes
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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