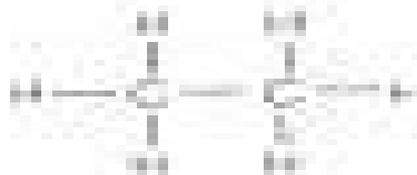


MICHEL 1: Relative electronegativities: $O > N > C > H$

A



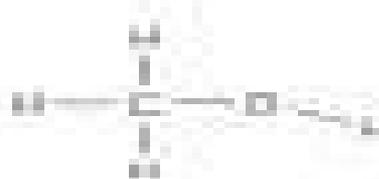
B



C



D



The bonds in B and C are polar.
 The molecules A and D are non-polar.

exercises

1. For each molecule indicated 1, draw in the polar covalent bond between the atoms. Write each element's valence electrons below:

- a. H₂ molecule in both electronegativity (the more electronegative element should be the more electronegative atom)
- b. H₂ both elements equally electronegative, draw the electron in the middle of the bond.

2. For each molecule, draw the distribution of the valence electrons that are shared completely. Draw circles or partial charge around the atoms. δ^+ or δ^-

3. Label all polar bonds with an arrow (\rightarrow).

4. Label the polar bond on a line to indicate the direction of the partial charge. Label the bond with an arrowhead for the direction. Draw the polarisation vector inside and pass through the C-Cl bond.

5. Draw molecules that are polar (have a net polar molecule) and non-polar molecules (overall)? Be particularly intelligent in drawing.

Biological Molecules Pogil Activities Answers Key Pdf

Arshad Iqbal



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