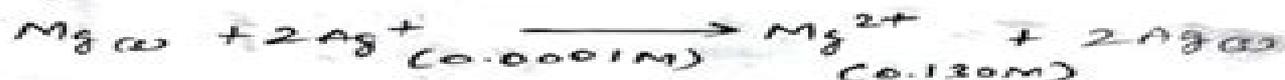


ELECTROCHEMISTRY NUMERICALS:

EXAMPLE PROBLEMS:

Prob: 3.1 Represent the cell in which the following reaction takes place



Calculate its E_{cell} if $E_{\text{cell}}^{\circ} = 3.17\text{V}$

Soln:

cell notation:



$$\begin{aligned} E_{\text{cell}} &= E_{\text{cell}}^{\circ} - \frac{0.0591\text{V}}{2} \log \frac{0.130}{(0.0001)^2} \\ &= 3.17\text{V} - 0.02955 \log \frac{0.13}{(10^{-4})^2} \\ &= 3.17 - 0.02955 \log \frac{0.13}{10^{-8}} \\ &= 3.17 - 0.02955 \log (0.13 \times 10^8) \\ &= 3.17 - 0.02955 \log (13 \times 10^6) \\ &= 3.17 - 0.02955 \log 13 + \log 10^6 \\ &= 3.17 - 0.02955 (1.1139 + 6) \\ &= 3.17 - (0.02955 \times 7.1139) \\ \therefore E_{\text{cell}} &= 3.17 - 0.21 = 2.96\text{V} \checkmark \end{aligned}$$

Electrochemistry Problems And Solutions

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Solved Problems in Electrochemistry for Universities and Industry Dominique L. Piron, 2004 **Electrochemical Methods: Fundamentals and Applications, 3e Student Solutions Manual** Cynthia G. Zoski, John A. Leddy, Allen J. Bard, Larry R. Faulkner, Henry S. White, 2024-11-04 Provides students with solutions to problems in the 3rd edition of the classic textbook *Electrochemical Methods Fundamentals and Applications*. *Electrochemical Methods* is a popular textbook on electrochemistry that takes the reader from the most basic chemical and physical principles through fundamentals of thermodynamics, kinetics, and mass transfer all the way to a thorough treatment of all important experimental methods. Holistically, it offers comprehensive coverage of all important topics in the field. To aid in reader comprehension, exercises are included at the end of each chapter which extend concepts introduced in the text or show how experimental data are reduced to fundamental results. This book provides worked solutions for many of the end of chapter exercises and is a key resource for any student who makes use of the original textbook.

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