

Oxidation number of IA group (+1)
Oxidation number of II-A group (+2)
Oxidation number of III-A group (+3)

Examples

$K_2Cr_2O_7$
 $(+1)2 + 2Cr - 2(7) = 0$
 $+2 + 2Cr - 14 = 0$
 $2Cr = +14 - 2 = +12$
 $Cr = +6$

$NaBrO_3$
 $1 + Br + (-2)(3) = 0$
 $1 + Br - 6 = 0$
 $Br - 5 = 0$
 $Br = +5$

ELECTROLYTIC CONDUCTION

Electrolytic Conduction	Electronic Conduction
<ul style="list-style-type: none"> Conduction through fused or aqueous electrolyte It is due to mobile ions Conduction increases with increase in temperature due to increase in ionization 	<ul style="list-style-type: none"> Conduction through solid metal Due to free electrons Conduction decreases with temperature due to increase in oscillations of cations

Products of electrolysis
 (when electrodes take part in the reaction)

Electrolyte	Copper cathode	Copper anode
$CuSO_4(aq)$	Cu deposits	$Cu_{(s)}$ dissolves to form Cu^{2+} ions
Electrolyte	Silver cathode	Silver anode
$AgNO_3(aq)$ and $HNO_3(aq)$	Ag deposits	$Ag_{(s)}$ dissolves to form Ag^+ ions

Products of electrolysis
 (using inert electrodes of platinum or graphite)

Electrolyte	Cathode	Anode
$PbBr_2$ (molten)	$Pb_{(s)}$	$Br_{2(g)}$
$NaCl$ (molten)	$Na_{(s)}$	$Cl_{2(g)}$
$NaCl_{(aq)}$	$H_2(g)$	$Cl_{2(g)}$
$CuCl_2(aq)$	$Cu_{(s)}$	$Cl_{2(g)}$
$CuSO_4(aq)$	$Cu_{(s)}$	$O_{2(g)}$
$KNO_3(aq)$	$H_2(g)$	$O_{2(g)}$
$NaOH(aq)$	$H_2(g)$	$O_{2(g)}$
$H_2SO_4(aq)$	$H_2(g)$	$O_{2(g)}$

USES OF ELECTROLYSIS

Particular	Description
Extraction of Na metal	Fused $NaCl$ is used Down's cell is an example
Caustic soda from brine	Nelson cell (aqueous $NaCl$ is used)
Mg and Ca extraction	Extraction occurs from their fused chlorides
Extraction of Al	Electrolysis of fused bauxite is conducted
Anodized Al	Prepared by making anode of Al
Purification of Cu	Anode is made up of impure copper
Electroplating	Copper, silver, nickel or chromium plating is done in electrolytic cell

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Table of Contents Electrochemistry Problems And Answers

1. Understanding the eBook Electrochemistry Problems And Answers
 - The Rise of Digital Reading Electrochemistry Problems And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Electrochemistry Problems And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electrochemistry Problems And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electrochemistry Problems And Answers
 - Personalized Recommendations
 - Electrochemistry Problems And Answers User Reviews and Ratings

- Electrochemistry Problems And Answers and Bestseller Lists
- 5. Accessing Electrochemistry Problems And Answers Free and Paid eBooks
 - Electrochemistry Problems And Answers Public Domain eBooks
 - Electrochemistry Problems And Answers eBook Subscription Services
 - Electrochemistry Problems And Answers Budget-Friendly Options
- 6. Navigating Electrochemistry Problems And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Electrochemistry Problems And Answers Compatibility with Devices
 - Electrochemistry Problems And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electrochemistry Problems And Answers
 - Highlighting and Note-Taking Electrochemistry Problems And Answers
 - Interactive Elements Electrochemistry Problems And Answers
- 8. Staying Engaged with Electrochemistry Problems And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electrochemistry Problems And Answers
- 9. Balancing eBooks and Physical Books Electrochemistry Problems And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electrochemistry Problems And Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electrochemistry Problems And Answers
 - Setting Reading Goals Electrochemistry Problems And Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electrochemistry Problems And Answers
 - Fact-Checking eBook Content of Electrochemistry Problems And Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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