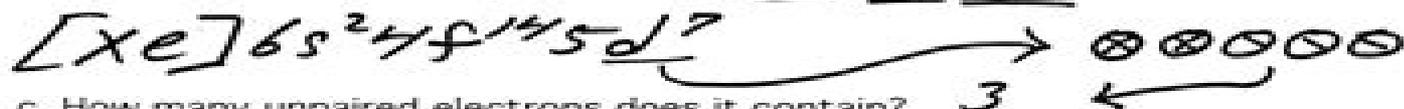


2. For Iridium-178 (atomic number 77)

a. Determine the number of protons, neutrons, and electrons

$$\#p = 77 \quad \#n = 101 \quad \#e = 77$$

b. Write its electron configuration using noble gas notation



c. How many unpaired electrons does it contain? 3

33. Name the element represented by each of the following electron configurations and determine the number of unpaired electrons each contains:



Sulfur

unpaired e⁻

2



Rhodium

2



Lead

2

34. Determine which of the following electron configurations represent an atom in the excited state and identify that element. (Circle choice and name the element)



35. Three isotopes of argon occur in nature, Ar-36, Ar-38, and Ar-40. Calculate the atomic mass of argon to two decimal places, given the following relative atomic masses and abundances of each of the isotopes: argon-36 (35.97 u; 0.337%), argon-38 (37.96 u; 0.063%), and argon-40 (39.96 u; 99.600%).

$$0.00337(35.97) + 0.00063(37.96) + 0.99600(39.96) = \boxed{39.95 \text{ u}}$$

36. Calculate the mass in grams of each of the following:

a. 3.00 mol Al

$$3.00 \text{ mol Al} \times \frac{26.98 \text{ g}}{1 \text{ mol}} = \boxed{80.9 \text{ g}}$$

b. 2.56×10^{24} atoms Li

$$2.56 \times 10^{24} \text{ atoms} \times \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{6.94 \text{ g}}{1 \text{ mol}} = \boxed{29.5 \text{ g Li}}$$

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M Mark



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Table of Contents Calculating Average Atomic Mass Worksheet Answers

1. Understanding the eBook Calculating Average Atomic Mass Worksheet Answers
 - The Rise of Digital Reading Calculating Average Atomic Mass Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Calculating Average Atomic Mass Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Calculating Average Atomic Mass Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Calculating Average Atomic Mass Worksheet Answers
 - Personalized Recommendations
 - Calculating Average Atomic Mass Worksheet Answers User Reviews and Ratings
 - Calculating Average Atomic Mass Worksheet Answers and Bestseller Lists

5. Accessing Calculating Average Atomic Mass Worksheet Answers Free and Paid eBooks
 - Calculating Average Atomic Mass Worksheet Answers Public Domain eBooks
 - Calculating Average Atomic Mass Worksheet Answers eBook Subscription Services
 - Calculating Average Atomic Mass Worksheet Answers Budget-Friendly Options
6. Navigating Calculating Average Atomic Mass Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Calculating Average Atomic Mass Worksheet Answers Compatibility with Devices
 - Calculating Average Atomic Mass Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Calculating Average Atomic Mass Worksheet Answers
 - Highlighting and Note-Taking Calculating Average Atomic Mass Worksheet Answers
 - Interactive Elements Calculating Average Atomic Mass Worksheet Answers
8. Staying Engaged with Calculating Average Atomic Mass Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Calculating Average Atomic Mass Worksheet Answers
9. Balancing eBooks and Physical Books Calculating Average Atomic Mass Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Calculating Average Atomic Mass Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Calculating Average Atomic Mass Worksheet Answers
 - Setting Reading Goals Calculating Average Atomic Mass Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Calculating Average Atomic Mass Worksheet Answers
 - Fact-Checking eBook Content of Calculating Average Atomic Mass Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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