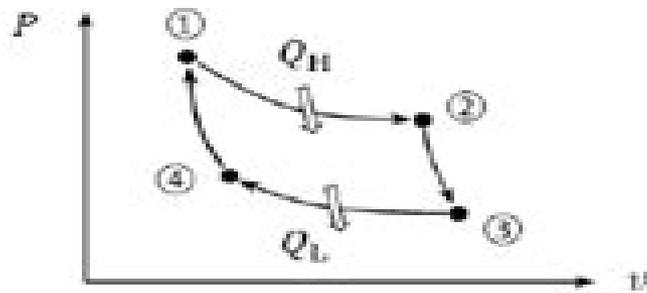


## Written Problem

The Carnot cycle is an idealization that involves four processes as shown (see also Figure 6–37 in textbook). As each of the processes is reversible, the entire cycle is also reversible.



- ① → ② : Isothermal expansion at  $T = T_H$
- ② → ③ : Adiabatic expansion (no heat transfer)
- ③ → ④ : Isothermal compression at  $T = T_L$
- ④ → ① : Adiabatic compression (no heat transfer)

- (a) Suppose the Carnot cycle is carried out in a closed system with an ideal gas as working fluid. Show that

$$Q_H = mRT_H \ln(V_2/V_1)$$

where  $m$  is the mass of the ideal gas. (Recall that the internal energy of an ideal gas depends *only* on the temperature  $T$ .)

You are now provided with more specifics: The ideal gas is 2 g of air. The Carnot cycle operates between  $T_H = 500$  K and  $T_L = 300$  K, with  $V_2 = 2V_1$ .

- (b) What is the amount of heat  $Q_L$  (kJ) that is transferred to the cold reservoir in every cycle?
- (c) Determine  $W_{\text{net}}$  (kJ), the net amount of work that is produced per cycle.

Note that  $W_{\text{net}}$  can be obtained by summing the  $\int P \, dV$  contribution from each of the four processes. This approach, however, is not recommended as there are much easier ways to get to the answer.

# Carnot Cycle Problems And Solutions

**Royal Society (Great Britain)**



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## **Table of Contents Carnot Cycle Problems And Solutions**

1. Understanding the eBook Carnot Cycle Problems And Solutions
  - The Rise of Digital Reading Carnot Cycle Problems And Solutions
  - Advantages of eBooks Over Traditional Books
2. Identifying Carnot Cycle Problems And Solutions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Carnot Cycle Problems And Solutions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Carnot Cycle Problems And Solutions
  - Personalized Recommendations
  - Carnot Cycle Problems And Solutions User Reviews and Ratings
  - Carnot Cycle Problems And Solutions and Bestseller Lists
5. Accessing Carnot Cycle Problems And Solutions Free and Paid eBooks

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  - Carnot Cycle Problems And Solutions eBook Subscription Services
  - Carnot Cycle Problems And Solutions Budget-Friendly Options
6. Navigating Carnot Cycle Problems And Solutions eBook Formats
    - ePub, PDF, MOBI, and More
    - Carnot Cycle Problems And Solutions Compatibility with Devices
    - Carnot Cycle Problems And Solutions Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Carnot Cycle Problems And Solutions
    - Highlighting and Note-Taking Carnot Cycle Problems And Solutions
    - Interactive Elements Carnot Cycle Problems And Solutions
  8. Staying Engaged with Carnot Cycle Problems And Solutions
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Carnot Cycle Problems And Solutions
  9. Balancing eBooks and Physical Books Carnot Cycle Problems And Solutions
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Carnot Cycle Problems And Solutions
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Carnot Cycle Problems And Solutions
    - Setting Reading Goals Carnot Cycle Problems And Solutions
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Carnot Cycle Problems And Solutions
    - Fact-Checking eBook Content of Carnot Cycle Problems And Solutions
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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