

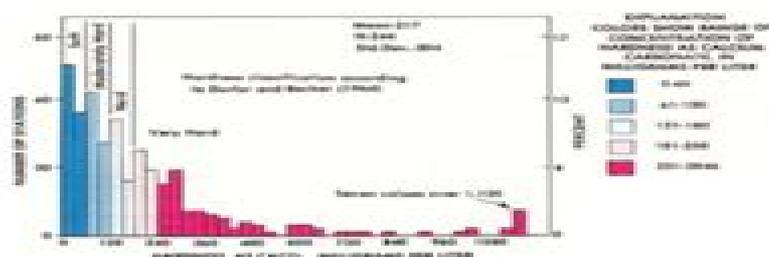
# EDTA Titration for Determination of Calcium and Magnesium

Before attempting this experiment, you may need to consult the section in the textbook dealing with complexometric titrations. In this procedure a stock solution of EDTA will be standardized against primary standard  $\text{CaCO}_3$ . This standardized EDTA solution is then used to determine water "hardness" – the total calcium and magnesium content, typically expressed as equivalent calcium carbonate mass concentration. While both metals are nutrients that are needed for good human health, high concentrations of these cations ( $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$ ) in water can cause deposits to form in bathrooms and kitchens, and cause the formation of "soap scum" – clogging pipes and decreasing the effectiveness of the cleanser. Therefore, very hard water is undesirable for use as tap water. Guidelines for classification of water as provided by the U.S. Geologic Survey are listed in the table below. The USGS website has more information about water hardness and alkalinity if you are interested. We're fortunate in that most of the water in the Portland area is soft.

**Table 1 U.S. Geological Survey Classification of Water Hardness**

(Source: <http://water.usgs.gov/wmq/hardness-alkalinity.html>)

$\text{CaCO}_3$ (mg/L)	Hardness
0-60	soft
61-120	moderately hard
121-180	hard
>180	very hard



**Figure 1 Hardness classification from Briggs et al. 1977**

Briggs, J.C., and Ficke, J.F., 1977, Quality of Rivers of the United States, 1975 Water Year -- Based on the National Stream Quality Accounting Network (NASQAN): U.S. Geological Survey Open-File Report 78-200, 436 p.

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## **Table of Contents Calcium Analysis By Edta Titration**

1. Understanding the eBook Calcium Analysis By Edta Titration
  - The Rise of Digital Reading Calcium Analysis By Edta Titration
  - Advantages of eBooks Over Traditional Books
2. Identifying Calcium Analysis By Edta Titration
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Calcium Analysis By Edta Titration
  - User-Friendly Interface
4. Exploring eBook Recommendations from Calcium Analysis By Edta Titration
  - Personalized Recommendations
  - Calcium Analysis By Edta Titration User Reviews and Ratings
  - Calcium Analysis By Edta Titration and Bestseller Lists
5. Accessing Calcium Analysis By Edta Titration Free and Paid eBooks
  - Calcium Analysis By Edta Titration Public Domain eBooks
  - Calcium Analysis By Edta Titration eBook Subscription Services

- Calcium Analysis By Edta Titration Budget-Friendly Options
- 6. Navigating Calcium Analysis By Edta Titration eBook Formats
  - ePub, PDF, MOBI, and More
  - Calcium Analysis By Edta Titration Compatibility with Devices
  - Calcium Analysis By Edta Titration Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Calcium Analysis By Edta Titration
  - Highlighting and Note-Taking Calcium Analysis By Edta Titration
  - Interactive Elements Calcium Analysis By Edta Titration
- 8. Staying Engaged with Calcium Analysis By Edta Titration
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Calcium Analysis By Edta Titration
- 9. Balancing eBooks and Physical Books Calcium Analysis By Edta Titration
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Calcium Analysis By Edta Titration
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Calcium Analysis By Edta Titration
  - Setting Reading Goals Calcium Analysis By Edta Titration
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Calcium Analysis By Edta Titration
  - Fact-Checking eBook Content of Calcium Analysis By Edta Titration
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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