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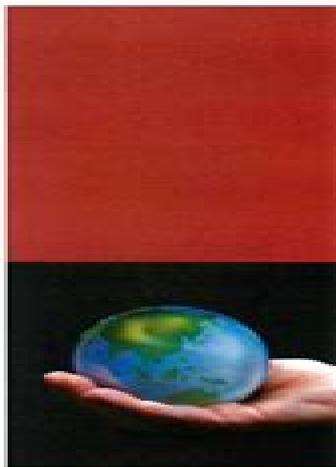
Business English Practice File

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David Cotton · David Falvey · Simon Kent



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time the total pressure $p_a \times \text{total mole fraction}$ a left frac mole a moles total right sum $n_1 \text{ all } x_{n_1}$ mole fraction of substance
a is often denoted by x_a and

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mass C_6H_6 6 12 6 1 g mol 1 78 g mol 1 hence C_6H_6 s number of moles $\frac{30}{78} \text{ mol} = 0.3846 \text{ mol}$ molar mass of carbon tet CCl_4 1
12 4 355 154 g mol 1

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g of water calculate the molarity and molality of the solution given $n_1 \text{ h } 1 \text{ c } 12 \text{ o } 16$ given mass of solute urea 11 11 g mass
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