

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and make use of Le Chatelier's analysis and an equilibrium shift construction through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction given by part (a).
2. Write a balanced equation for the reaction given by part (b).
3. Write a balanced equation for the reaction given by part (c).
4. In part (a), what is the conjugate acid of the base HCO_3^- and what is the conjugate base of the acid H_2CO_3 ?
Compare with those in the second group and show the other acids and bases. The conjugate acid of a base is the species that formed in the process and the conjugate base of an acid is the species that remains.
5. What is the conjugate acid of the base HCO_3^- in part (b)?
What is the conjugate base of the acid H_2CO_3 in part (b)?
6. What is the conjugate acid of the base HCO_3^- in part (c)?
What is the conjugate base of the acid H_2CO_3 in part (c)?
7. Write the equation for the reaction between H^+ and OH^- ions:
 $\text{H}^+ + \text{OH}^- \rightleftharpoons \text{H}_2\text{O}$
8. Write the equation for the reaction between H^+ and CO_3^{2-} ions:
 $\text{H}^+ + \text{CO}_3^{2-} \rightleftharpoons \text{HCO}_3^-$
9. Write the equation for the reaction between H^+ and HCO_3^- ions:
 $\text{H}^+ + \text{HCO}_3^- \rightleftharpoons \text{H}_2\text{CO}_3$
10. Write the equation for the reaction between H^+ and H_2CO_3 ions:
 $\text{H}^+ + \text{H}_2\text{CO}_3 \rightleftharpoons \text{H}_3\text{CO}_3^+$

Investigating Chemical Equilibrium Lab 12a Answer

Wenbin Ji



Investigating Chemical Equilibrium Lab 12a Answer:

Ignite the flame of optimism with Crafted by is motivational masterpiece, Fuel Your Spirit with **Investigating Chemical Equilibrium Lab 12a Answer** . In a downloadable PDF format (*), this ebook is a beacon of encouragement. Download now and let the words propel you towards a brighter, more motivated tomorrow.

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