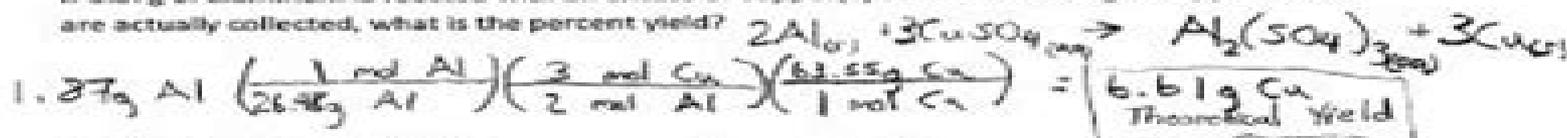


Stoichiometry Worksheet #2

1. If 1.87g of aluminum is reacted with an excess of copper(II) sulfate and 3.74g of copper metal are actually collected, what is the percent yield?



$$\frac{3.74g \text{ Cu (Actual Yield)}}{6.61 \text{ (Theoretical)}} \times 100 = \boxed{56.6\% \text{ Yield}}$$

2. 0.531 mol of iron(III) oxide are reacted with an excess of carbon monoxide and 57.8g of iron is produced. What is the percent yield? Hint: The other product of the reaction is carbon dioxide.



$$\frac{57.8g \text{ (Actual)}}{59.3g \text{ (Theor)}} \times 100 = \boxed{97.5\% \text{ Yield}}$$

3. Calcium carbonate can be decomposed by heating to produce solid calcium oxide and carbon dioxide gas. If you wanted to produce 5.5L of carbon dioxide gas, what mass of calcium carbonate would have to be used? Assume 100% yield and conditions at STP.



$$CaCO_3 \rightarrow \begin{matrix} 40.08 \\ 12.01 \\ 3(16.00) \end{matrix}$$

$$5.5L \text{ CO}_2 \left( \frac{1 \text{ mol CO}_2}{22.4L \text{ CO}_2} \right) \left( \frac{1 \text{ mol CaCO}_3}{1 \text{ mol CO}_2} \right) \left( \frac{100.09g \text{ CaCO}_3}{1 \text{ mol CaCO}_3} \right) = \boxed{24.8g \text{ CaCO}_3}$$

4. Aluminum chloride, an inexpensive reagent used in many industrial processes, is made by reacting scrap aluminum with chlorine. If 2.70g of aluminum are reacted with 4.05g of chlorine, which is the limiting reagent and how many grams of the excess reagent remain when the reaction is complete?



$$.100 \text{ mol Al} - .037 \text{ mol Al} = .063 \text{ mol Al (Actual Product)}$$

$$.063 \text{ mol Al} \left( \frac{26.98g \text{ Al}}{1 \text{ mol Al}} \right) = 1.67g \text{ of Al in Excess}$$

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**Robert Nakon**



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## **Table of Contents Stoichiometry 8b Extra Practice Problems Answers**

1. Understanding the eBook Stoichiometry 8b Extra Practice Problems Answers
  - The Rise of Digital Reading Stoichiometry 8b Extra Practice Problems Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry 8b Extra Practice Problems Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Stoichiometry 8b Extra Practice Problems Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry 8b Extra Practice Problems Answers
  - Personalized Recommendations
  - Stoichiometry 8b Extra Practice Problems Answers User Reviews and Ratings
  - Stoichiometry 8b Extra Practice Problems Answers and Bestseller Lists

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  - Stoichiometry 8b Extra Practice Problems Answers Public Domain eBooks
  - Stoichiometry 8b Extra Practice Problems Answers eBook Subscription Services
  - Stoichiometry 8b Extra Practice Problems Answers Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
  - Stoichiometry 8b Extra Practice Problems Answers Compatibility with Devices
  - Stoichiometry 8b Extra Practice Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Stoichiometry 8b Extra Practice Problems Answers
  - Highlighting and Note-Taking Stoichiometry 8b Extra Practice Problems Answers
  - Interactive Elements Stoichiometry 8b Extra Practice Problems Answers
8. Staying Engaged with Stoichiometry 8b Extra Practice Problems Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Stoichiometry 8b Extra Practice Problems Answers
9. Balancing eBooks and Physical Books Stoichiometry 8b Extra Practice Problems Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Stoichiometry 8b Extra Practice Problems Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry 8b Extra Practice Problems Answers
  - Setting Reading Goals Stoichiometry 8b Extra Practice Problems Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry 8b Extra Practice Problems Answers
  - Fact-Checking eBook Content of Stoichiometry 8b Extra Practice Problems Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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web ein historischer kriminalroman der im irland des 7 jahrhunderts spielt als eine frau von der herkunft und bildung schwester fidelmas noch über macht und einfluss verfügte

