

LIMITING REAGENT Practice Problems

1. At high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide:



In one experiment, 7.62 g of Fe are allowed to react with 8.67 g of S.

- What is the limiting reagent, and what is the reactant in excess?
 - Calculate the mass of FeS formed.
2. Acrylonitrile, $\text{C}_3\text{H}_3\text{N}$, is the starting material for the production of a kind of synthetic fiber (acrylics) and can be made from propylene, C_3H_6 , by reaction with nitric oxide, NO, as follows:



What mass of $\text{C}_3\text{H}_3\text{N}$ can be made when 21.6 g of C_3H_6 react with 21.6 g of nitric oxide?

3. Calculate the percent yield for the reaction: $\text{P}_4 \text{ (s)} + 6 \text{Cl}_2 \text{ (g)} \rightarrow 4 \text{PCl}_3 \text{ (l)}$
if 75.0 g of phosphorus reacts with excess chlorine gas to produce 111.0 g of phosphorus trichloride.

4. Formic acid, HCHO_2 , burns in oxygen to form carbon dioxide and water as follows:



If a 3.15-g sample of formic acid was burned in 2.0 L of oxygen, what volume of carbon dioxide would be produced? (Assume the reaction occurs at standard temperature and pressure, STP.)

5. Zinc metal reacts with hydrochloric acid to produce zinc chloride and hydrogen gas.
- Balance the following reaction: $\text{Zn (s)} + \text{HCl (aq)} \rightarrow \text{ZnCl}_2 \text{ (aq)} + \text{H}_2 \text{ (g)}$
 - A 3.50-g sample of zinc metal is allowed to react with 2.50 g of hydrochloric acid. Complete the following table:

Reactants/products	Zn (grams)	HCl (grams)	ZnCl ₂ (grams)	H ₂ (L)
Before reaction				
After reaction	1.26 g			

6. Consider the reaction: $\text{MnO}_2 + 4 \text{HCl} \rightarrow \text{MnCl}_2 + \text{Cl}_2 + 2 \text{H}_2\text{O}$
If 0.45 mols of MnO_2 can react with 48.2 g of HCl, how many grams of Cl_2 could be produced?

7. One of the components of the fuel mixture on the Apollo lunar module involved a reaction with hydrazine, N_2H_4 , and dinitrogen tetraoxide, N_2O_4 . If the balanced equation for this reaction is



what volume of N_2 gas (measured at STP) would result from the reaction of 1500 kg of hydrazine and 1000 kg of N_2O_4 ?

Stoichiometry Limiting Reagent Problems Answers

Katrin Zwirgmaier



Stoichiometry Limiting Reagent Problems Answers:

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Table of Contents Stoichiometry Limiting Reagent Problems Answers

1. Understanding the eBook Stoichiometry Limiting Reagent Problems Answers
 - The Rise of Digital Reading Stoichiometry Limiting Reagent Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Limiting Reagent Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Limiting Reagent Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Limiting Reagent Problems Answers
 - Personalized Recommendations
 - Stoichiometry Limiting Reagent Problems Answers User Reviews and Ratings
 - Stoichiometry Limiting Reagent Problems Answers and Bestseller Lists
5. Accessing Stoichiometry Limiting Reagent Problems Answers Free and Paid eBooks
 - Stoichiometry Limiting Reagent Problems Answers Public Domain eBooks
 - Stoichiometry Limiting Reagent Problems Answers eBook Subscription Services
 - Stoichiometry Limiting Reagent Problems Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Stoichiometry Limiting Reagent Problems Answers Compatibility with Devices
 - Stoichiometry Limiting Reagent Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Limiting Reagent Problems Answers
 - Highlighting and Note-Taking Stoichiometry Limiting Reagent Problems Answers
 - Interactive Elements Stoichiometry Limiting Reagent Problems Answers

8. Staying Engaged with Stoichiometry Limiting Reagent Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Limiting Reagent Problems Answers
9. Balancing eBooks and Physical Books Stoichiometry Limiting Reagent Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Limiting Reagent Problems Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Limiting Reagent Problems Answers
 - Setting Reading Goals Stoichiometry Limiting Reagent Problems Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Limiting Reagent Problems Answers
 - Fact-Checking eBook Content of Stoichiometry Limiting Reagent Problems Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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