

pH Properties of Buffer Solutions

Flinn AP Chemistry Laboratory #19

Introduction

One of the most important applications of acids and bases in chemistry and biology is that of buffers. A buffer solution resists rapid changes in pH when acids and bases are added to it. Every living cell contains natural buffer systems to maintain the constant pH needed for proper cell function. Many consumer products are also buffered to safeguard their activity. What are buffers made of? How do buffers maintain the delicate pH balance needed for life and health?

Concepts

- Buffers
- Weak acid-weak base
- Conjugate acid-base pairs
- Dissociation constant
- Ideal buffer
- Background

The ability of buffers to resist changes in pH when acid or base is added is a result of their chemical composition. All buffers contain a mixture of a conjugate acid-base pair; either a weak acid (HA) and its conjugate base (A⁻), or a weak base (B), and its conjugate acid (BH⁺). Weak acids and bases both dissociate slightly in water (Reactions 1 and 2).



weak acid *conjugate base*



weak base *conjugate acid*

These reactions are reversible and both the weak acid and its conjugate base or the weak base and its conjugate acid are present in solution.

The equilibrium constant expression for these dissociation reactions are:

$$K_a = [\text{H}_3\text{O}^+][\text{A}^-]/[\text{HA}] \quad \text{Equation 1}$$

$$K_b = [\text{OH}^-][\text{BH}^+]/[\text{B}] \quad \text{Equation 2}$$

Buffers control pH because the two buffering compounds, either HA and A⁻ or B and BH⁺, are able to neutralize both acids and bases added to the solution.



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Table of Contents Properties Of Buffer Solutions Lab 16

1. Understanding the eBook Properties Of Buffer Solutions Lab 16
 - The Rise of Digital Reading Properties Of Buffer Solutions Lab 16
 - Advantages of eBooks Over Traditional Books
2. Identifying Properties Of Buffer Solutions Lab 16
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Properties Of Buffer Solutions Lab 16
 - User-Friendly Interface
4. Exploring eBook Recommendations from Properties Of Buffer Solutions Lab 16
 - Personalized Recommendations
 - Properties Of Buffer Solutions Lab 16 User Reviews and Ratings
 - Properties Of Buffer Solutions Lab 16 and Bestseller Lists
5. Accessing Properties Of Buffer Solutions Lab 16 Free and Paid eBooks
 - Properties Of Buffer Solutions Lab 16 Public Domain eBooks
 - Properties Of Buffer Solutions Lab 16 eBook Subscription Services
 - Properties Of Buffer Solutions Lab 16 Budget-Friendly Options
6. Navigating Properties Of Buffer Solutions Lab 16 eBook Formats

- ePub, PDF, MOBI, and More
 - Properties Of Buffer Solutions Lab 16 Compatibility with Devices
 - Properties Of Buffer Solutions Lab 16 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Properties Of Buffer Solutions Lab 16
 - Highlighting and Note-Taking Properties Of Buffer Solutions Lab 16
 - Interactive Elements Properties Of Buffer Solutions Lab 16
 8. Staying Engaged with Properties Of Buffer Solutions Lab 16
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Properties Of Buffer Solutions Lab 16
 9. Balancing eBooks and Physical Books Properties Of Buffer Solutions Lab 16
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Properties Of Buffer Solutions Lab 16
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Properties Of Buffer Solutions Lab 16
 - Setting Reading Goals Properties Of Buffer Solutions Lab 16
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Properties Of Buffer Solutions Lab 16
 - Fact-Checking eBook Content of Properties Of Buffer Solutions Lab 16
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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