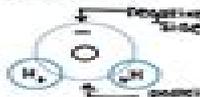


Structure of H₂O
hydrogen bonding
Denser at 4°C



unit one guide

Biological molecules
proteins
nucleic acids

H₂O provides cushion, adhesion, surface tension, high specific heat, high heat of vaporization, variable density, greater density as a solid.

THIS GENE: when ions → water expands due to hydrogen bonding, molecules next to each other.

hydrogen bonds require lots of energy to change state of matter

CONDENSATION → when two molecules produce hydrogen and hydroxyl ions, releases H₂O

Hydroxyl → attracted to same surface
Hydrogen bond → attraction between surfaces of two liquid molecules

C	H	N	O	P	S
Carbon	Hydrogen	Nitrogen	Oxygen	Phosphorus	Sulfur

more hydrogen present → C⁻ bond, biological molecules, cell formation
N⁺ bond proteins, nucleic acids
P⁻ bond nucleic acids + lipids
→ make up 98% of organisms

Carbohydrates ↓ C, H, O	Lipids ↓ C, H, O	Nucleic acids ↓ C, H, N, O, P	Proteins ↓ C, H, N, O, S
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Condensation synthesis → the removal of an H₂O molecule to bond monomers

Nucleic acid structure:

• One carbon sugar (deoxyribose/ribose), phosphate, nitrogenous base (A, T, C, G), exact copies of strands



Starch: 1,4 glucose
Glycogen: 1,4 glucose
Cellulose: 1,4 glucose
Chitin: 1,4 glucose

Proteins:

Specific order of amino acids → primary structure

R group interactions determine structure + function



PROTEINS

• amino acids bond via a peptide bond (covalent) through dehydration synthesis
• R group change can cause a hydrophobic protein to lose the place of a hydrophilic protein

membranes

• made up of phospholipids
• create selectively permeable barrier
• contain phospholipids, cholesterol, and proteins



LIPIDS

Phospholipid macromolecules

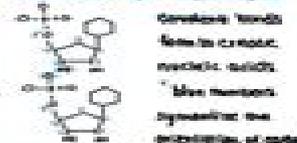
• differences in saturation determine structure and function
• Cells, sterols, phospholipids
• Saturated → no double bonds
• Unsaturated → has double bonds
• Lipids: triglycerides, cholesterol, most of the membrane



Amino Acids:

• have directionality
• linear chains, connected by covalent bonds
• amino end and carboxyl end

DEHYDRATION SYNTHESIS



covalent bonds form to create nucleic acids
• the number specifies the structure of water

PROTEIN STRUCTURE

• Primary: sequence of amino acids, specific order determine protein shape
• Secondary: how folding of chain into alpha helix and beta sheet occurs
• Tertiary: 3-D shape, hydrogen bonds, disulfide link, ionic bonds, hydrophobic int.
• Quaternary: interactions between multiple polypeptide chains

+ FUNCTION

• Enzymes → speeds up reaction rates in the cell
• Cell structure → part of the cellular membrane
• Communication between cells

DENATURATION:

• the protein starts to lose its shape/structure when (bonds break)
• Temperature change
• pH of surrounding area
• Salt concentration

NH₂ = Amino
COOH = Carboxyl

NUCLEIC ACIDS

• add new nucleotides to the 3' end
• base nitrogenous bases: purine + pyrimidine
• pyrimidine is single ringed, purine is double
• Antiparallel double helix
• 5' on phosphate end, 3' on hydroxyl end
• both back and forth have nitrogenous bases (purines) + sugar-phosphate backbone

ANTIPARALLEL



Chaperones:

• protein molecules that assist protein folding

ALDOSE VS KETOSE

• contains aldehyde functional group

ALPHA VS BETA

• contains ketone functional group



ALPHA VS BETA

• all monomers are in the same orientation



ALPHA VS BETA

• Every glucose monomer is upside down from its neighbor

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H Kauffman



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