

Practicing with Isotopes

1. Complete the following questions. Assume all the atoms are neutral.



Element: _____

of protons: _____

of neutrons: _____



Element: _____

of protons: _____

of neutrons: _____



Element: _____

of protons: _____

of neutrons: _____



Element: _____

of protons: _____

of neutrons: _____



Element: _____

of protons: _____

of neutrons: _____



Element: _____

of protons: _____

of neutrons: _____



Element: _____

of protons: _____

of neutrons: _____



Element: _____

of protons: _____

of neutrons: _____



Element: _____

of protons: _____

of neutrons: _____

2. Calculate the average atomic mass for the following element from the given information about the relevant isotopes:

Isotope 1: Mass = 35 amu

Abundance = 75.53%

Isotope 2: Mass = 37 amu

Abundance = 24.47%

Which element is this? _____

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Robert R. Redfield



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